

FORM PTO-1390
(REV 10-2000)

U.S. DEPARTMENT OF COMMERCE PATENT AND TRADEMARK OFFICE

ATTORNEY'S DOCKET NUMBER

9793822-0118

U.S. APPLICATION NO. (If known, see 37 CFR 1.5)

09/701950

TRANSMITTAL LETTER TO THE UNITED STATES
DESIGNATED/ELECTED OFFICE (DO/EO/US)
CONCERNING A FILING UNDER 35 U.S.C. 371INTERNATIONAL APPLICATION NO.
PCT/JP00/01915INTERNATIONAL FILING DATE
28 March 2000PRIORITY DATE CLAIMED
06 April 1999TITLE OF INVENTION Method for Producing Positive Electrode Active Material for Producing
Non-Aqueous Electrolyte Secondary BatteryAPPLICANT(S) FOR DO/EO/US
LI, Guohua; YAMADA, Atsuo; AZUMA, Hideto

Applicant herewith submits to the United States Designated/Elected Office (DO/EO/US) the following items and other information:

1. This is a **FIRST** submission of items concerning a filing under 35 U.S.C. 371.
2. This is a **SECOND** or **SUBSEQUENT** submission of items concerning a filing under 35 U.S.C. 371.
3. This is an express request to promptly begin national examination procedures (35 U.S.C. 371(f)).
4. The US has been elected by the expiration of 19 months from the priority date (PCT Article 31).
5. A copy of the International Application as filed (35 U.S.C. 371(c)(2))
 - a. is attached hereto (required only if not communicated by the International Bureau).
 - b. has been communicated by the International Bureau.
 - c. is not required, as the application was filed in the United States Receiving Office (RO/US).
6. An English language translation of the International Application as filed (35 U.S.C. 371(c)(2)).
7. Amendments to the claims of the International Application under PCT Article 19 (35 U.S.C. 371(c)(3))
 - a. are attached hereto (required only if not communicated by the International Bureau).
 - b. have been communicated by the International Bureau.
 - c. have not been made; however, the time limit for making such amendments has NOT expired.
 - d. have not been made and will not be made.
8. An English language translation of the amendments to the claims under PCT Article 19 (35 U.S.C. 371(c)(3)).
9. An oath or declaration of the inventor(s) (35 U.S.C. 371(c)(4)).
10. An English language translation of the annexes to the International Preliminary Examination Report under PCT Article 36 (35 U.S.C. 371(c)(5)).

Items 11 to 16 below concern document(s) or information included:

11. An Information Disclosure Statement under 37 CFR 1.97 and 1.98.
12. An assignment document for recording. A separate cover sheet in compliance with 37 CFR 3.28 and 3.31 is included.
13. A **FIRST** preliminary amendment.
- A **SECOND** or **SUBSEQUENT** preliminary amendment.
14. A substitute specification.
15. A change of power of attorney and/or address letter.
16. Other items or information:
Drawings (31 sheets - Figs. 1-31)
Express Mail

U.S. APPLICATION NO. (If known, see 37 CFR 1.5)

097701950

INTERNATIONAL APPLICATION NO
PCT/JP00/01915ATTORNEY'S DOCKET NUMBER
9793822-011817. The following fees are submitted:

BASIC NATIONAL FEE (37 CFR 1.492 (a) (1) - (5)):

Neither international preliminary examination fee (37 CFR 1.482) nor international search fee (37 CFR 1.445(a)(2)) paid to USPTO and International Search Report not prepared by the EPO or JPO \$1000.00

International preliminary examination fee (37 CFR 1.482) not paid to USPTO but International Search Report prepared by the EPO or JPO \$860.00

International preliminary examination fee (37 CFR 1.482) not paid to USPTO but international search fee (37 CFR 1.445(a)(2)) paid to USPTO \$710.00

International preliminary examination fee paid to USPTO (37 CFR 1.482) but all claims did not satisfy provisions of PCT Article 33(1)-(4) \$690.00

International preliminary examination fee paid to USPTO (37 CFR 1.482) and all claims satisfied provisions of PCT Article 33(1)-(4) \$100.00

ENTER APPROPRIATE BASIC FEE AMOUNT =

\$ 860.00

Surcharge of \$130.00 for furnishing the oath or declaration later than 20 30 months from the earliest claimed priority date (37 CFR 1.492(e)).

\$ 130.00

CLAIMS	NUMBER FILED	NUMBER EXTRA	RATE	
Total claims	26 - 20 =	6	X \$18.00	\$ 108.00
Independent claims	6 - 3 =	3	X \$80.00	\$ 240.00
MULTIPLE DEPENDENT CLAIM(S) (if applicable)			+ \$270.00	\$ 0
			TOTAL OF ABOVE CALCULATIONS =	\$ 1,338.00

<input checked="" type="checkbox"/> Applicant claims small entity status. See 37 CFR 1.27. The fees indicated above are reduced by 1/2.	\$ 0	

<input checked="" type="checkbox"/> SUBTOTAL =	\$ 1,338.00	
Processing fee of \$130.00 for furnishing the English translation later than <input type="checkbox"/> 20 <input type="checkbox"/> 30 months from the earliest claimed priority date (37 CFR 1.492(f)).	\$ 0	

<input checked="" type="checkbox"/> TOTAL NATIONAL FEE =	\$ 1,338.00	
Fee for recording the enclosed assignment (37 CFR 1.21(h)). The assignment must be accompanied by an appropriate cover sheet (37 CFR 3.28, 3.31). \$40.00 per property	+ \$ 0	

<input checked="" type="checkbox"/> TOTAL FEES ENCLOSED =	\$ 1,338.00	
	Amount to be refunded:	\$
	charged:	\$

a. A check in the amount of \$ 1,338.00 to cover the above fees is enclosed.

b. Please charge my Deposit Account No. _____ in the amount of \$_____ to cover the above fees. A duplicate copy of this sheet is enclosed.

c. The Commissioner is hereby authorized to charge any additional fees which may be required, or credit any overpayment to Deposit Account No. 19-3140. A duplicate copy of this sheet is enclosed.

NOTE: Where an appropriate time limit under 37 CFR 1.494 or 1.495 has not been met, a petition to revive (37 CFR 1.137(a) or (b)) must be filed and granted to restore the application to pending status.

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42,748
REGISTRATION NUMBER

31/phb

09/701950
526 Rec'd PCT/PTO 05 DEC 2000

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DESCRIPTION

Method for Producing Positive Electrode Active Material and Method for Producing Non-Aqueous Electrolyte Secondary Battery

Technical Field

This invention relates to a method for producing a positive electrode active material that is capable of reversibly doping/undoping lithium, and a method for producing a non-aqueous electrolyte secondary battery employing this positive electrode active material.

Background Art

Recently, with the marked progress in a variety of electronic equipment, researches in a rechargeable secondary battery, as a battery that can be used conveniently and economically for prolonged time, are underway. Typical of the known secondary batteries are a lead battery, an alkali storage battery and a lithium secondary battery.

Of these secondary batteries, a lithium secondary battery has advantages as to high output and high energy density. The lithium secondary battery is made up at least of positive and negative electrodes, containing active materials capable of reversibly introducing and removing lithium ions, and a non-aqueous electrolyte.

Currently, LiCoCO₂ is widely exploited as a positive electrode active material

of a lithium ion secondary battery having a potential of 4V with respect to the lithium potential. This LiCoCO_2 is of a high energy density and a high voltage and is an ideal positive electrode material in many respects. However, Co is localized in distribution and represents a rare resources, with the result that, if LiCoCO_2 is used as a positive electrode active material, the cost is raised, whilst stable supply is difficult.

Therefore, development of a electrode active material based on Ni, Mn or Fe, that is abundant in supply and inexpensive, is desirable. For example, LiNiO_2 , based on Ni, has a large theoretical capacity and a high discharging potential. However, in a battery employing LiNiO_2 , the LiNiO_2 crystal structure collapses with the charging/discharging cycles, so that the discharging capacity is lowered. On the other hand, LiNiO_2 suffers a drawback or poor thermal stability.

As an Mn-based electrode active material, there is proposed LiMn_2O_4 having a positive spinel structure and a spatial set $\text{Fd}3\text{m}$. This LiMn_2O_4 has a high potential of the order of 4V, with respect to the lithium potential, which is equivalent to that of LiCoCO_2 . Moreover, the LiMn_2O_4 is a highly promising material since it is easy to synthesize and has a high battery capacity. However, the battery constructed using LiMn_2O_4 is deteriorated in battery capacity since Mn is dissolved in the electrolytic solution with the charging/discharging cycles so that it is insufficient in stability or cycle characteristics.

On the other hand, such a material having LiFeO_2 as a basic structure is being researched as an Fe-based electrode active material. Although LiFeO_2 has a structure

similar to that of LiCoCO_2 or LiMn_2O_4 , it is unstable in structure and difficult to synthesize.

On the other hand, a compound having an olivinic structure as a positive electrode active material of the lithium secondary battery, such as, for example, a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$, where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M contains at least one of 3d transition metals (Fe, Mn, Co and Ni), is retained to be a promising material.

It is proposed in Japanese Laying-Open Patent H-9-171827 to use e.g., LiFePO_4 , amongst the compounds represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$, as a positive electrode of a lithium ion battery. This LiFePO_4 has a theoretical capacity as large as 170 mAh/g and contains one Li atom, that can be electro-chemically doped/undoped in an initial state, per Fe atom, and hence is a promising material as a positive electrode active material for the lithium ion battery.

Conventionally, this LiFePO_4 has been synthesized by sintering at a higher temperature of 800°C , under a reducing environment, using a bivalent iron salt, such as iron phosphate $\text{Fe}_3(\text{PO}_4)_2$ or iron acetate $\text{Fe}(\text{CH}_3\text{COO})_2$, as an Fe source which proves a starting material for synthesis.

However, Fe^{2+} is sensitive to a trace amount of oxygen contained in a synthesizing atmosphere and is readily oxidized to Fe^{3+} . The result is that trivalent iron compounds tend to co-exist in the produced $\text{Fe}_3(\text{PO}_4)_2$ to render it difficult to acquire single-phase LiFePO_4 .

It is reported in the above Publication that an actual battery fabricated using the $\text{Fe}_3(\text{PO}_4)_2$ synthesized by the above-described synthetic method has an actual capacity only as low as approximately 60 mAh/g to 70 mAh/g. Although the actual capacity of the order of 120 mAh/g is subsequently reported in the Journal of the Electrochemical Society, 144,1188 (1997), it cannot be said that a sufficient capacity has been achieved, in consideration that the theoretical capacity is 170 mAh/g.

If is LiFePO_4 is compared to LiMn_2O_4 , the former has a volumetric density of 3.6 g/cm³ and an average voltage of 3.4 V, whereas latter has a volumetric density of 4.2 g/cm³ and an average voltage of 3.9 V, with the capacity being 120 mAh/g. So, LiFePO_4 is lower than LiMn_2O_4 by approximately 10% in both the voltage and volumetric density. So, for the same capacity of 120 mAh/g, LiFePO_4 is lower than LiMn_2O_4 by not less than 10% and by not less than 20% in weight energy density and in volumetric energy density, respectively. Thus, in order to realize the energy density of LiFePO_4 which is of the level equivalent to or higher than LiMn_2O_4 , the capacity of 140 mAh/g or higher is required. However, this high capacity has not been realized with LiFePO_4 .

On the other hand, LiFePO_4 obtained by the conventional synthesizing method is lower in electrical conductivity than the positive electrode active material, such as LiMn_2O_4 , and hence suffers the problem of low load characteristics and poor molding performance. So, in order to prepare an electrode using LiFePO_4 as the positive electrode active material, a large quantity of electrically conductive material, such as

acetylene black, needs to be added, with the result that a non-aqueous electrolyte secondary battery employing LiFePO_4 as a positive electrode active material suffers a drawback that the energy density of the electrode is smaller, with the electrode molding performance being also lowered.

Disclosure of the Invention

It is an object of the present invention to provide a method for producing a positive electrode active material in which mixing of an oxidized 3d transition metal into a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ to realize a high capacity and a method for producing a high-capacity non-aqueous electrolyte secondary battery employing this positive electrode active material. It is a particular object of the present invention to provide a method for producing a positive electrode active material in which mixing of an oxidized 3d transition metal into a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ and a method for producing a high-capacity non-aqueous electrolyte secondary battery by employing this positive electrode active material to realize a high capacity in case the compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ is LiFePO_4 .

For accomplishing the above object, the present invention provides a method for producing a positive electrode active material including a mixing step of mixing a plurality of substances to give a precursor, the substances proving a starting material for synthesis of a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ where x is

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such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M includes at least one of 3d transition metals, and a sintering step of sintering and reacting the precursor obtained by the mixing step, wherein a reducing agent is added to the precursor in the mixing step.

In the method for producing the positive electrode active material according to the present invention, in which the reducing agent is added to the precursor in the mixing step, it is possible to prevent the 3d transition metal M in the starting material for synthesis from being oxidized by the residual oxygen to yield single-phase $Li_xM_yPO_4$ free of impurities.

The present invention also provides a method for producing a non-aqueous electrolyte secondary battery having a positive electrode containing a positive electrode active material capable of reversibly doping/undoping lithium, a negative electrode mounted facing the positive electrode and capable of reversibly doping/undoping lithium, and a non-aqueous electrolyte interposed between the positive electrode and the negative electrode, in which the positive electrode active material is produced by a method including a mixing step of mixing a plurality of substances, proving a starting material for synthesis of a compound represented by the general formula $Li_xM_yPO_4$ where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M includes at least one of 3d transition metals, to give a precursor, and a sintering step of sintering and reacting the precursor obtained by the mixing step, wherein a reducing agent is added to the precursor in the mixing step to prepare the positive

electrode active material.

In the method for producing the non-aqueous electrolyte secondary battery according to the present invention, in which the reducing agent is added to the precursor in the mixing step in the preparation of the positive electrode active material, it is possible to prevent the 3d transition metal M in the starting material for synthesis from being oxidized by the residual oxygen to yield single-phase $\text{Li}_x\text{M}_y\text{PO}_4$ free of impurities.

The present invention also provides a method for producing a positive electrode active material including a mixing step of mixing a plurality of substances proving a starting material for synthesis of a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M includes at least one of 3d transition metals, to give a precursor, a de-aerating step of removing air contained in the precursor obtained in the mixing step, a sintering step of sintering and reacting the precursor obtained by the mixing step.

In the method for producing the positive electrode active material according to the present invention, in which air contained in the precursor in the de-aerating step is removed in the de-aerating step, it is possible to prevent the 3d transition metal M in the starting material for synthesis from being oxidized by the residual oxygen to yield single-phase $\text{Li}_x\text{M}_y\text{PO}_4$ free of impurities.

The present invention also provides a method for producing a non-aqueous electrolyte secondary battery having a positive electrode containing a positive

electrode active material capable of reversibly doping/undoping lithium, a negative electrode mounted facing the positive electrode and capable of reversibly doping/undoping lithium, and a non-aqueous electrolyte interposed between the positive electrode and the negative electrode, in which the positive electrode active material is produced by a method including a mixing step of mixing a plurality of substances, proving a starting material for synthesis of a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M includes at least one of 3d transition metals, to give a precursor, a de-aerating step of de-aerating air contained in the precursor obtained in the mixing step and a sintering step of sintering and reacting the precursor obtained in a state free of air by the de-aerating step.

In the method for producing a non-aqueous electrolyte secondary battery according to the present invention, in which, in preparing the positive electrode active material, any air contained in the precursor is removed in the deaerating process in the preparation of the positive electrode active material, it is possible to prevent the 3d transition metal M in the starting material for synthesis from being oxidized by the residual oxygen to yield single-phase $\text{Li}_x\text{M}_y\text{PO}_4$ free of impurities.

It is another object of the present invention to provide a method for producing a positive electrode active material superior in load characteristics and molding performance, and which is capable of realizing a high capacity, and a method for producing a non-aqueous electrolyte secondary battery which, through the use of this

positive electrode active material, has a high capacity approaching to the theoretical capacity of 170 mAh/g.

For accomplishing this object, the present invention provides a method for producing a positive electrode active material including a mixing step of mixing a plurality of substances to give a precursor, the substances proving a starting material for synthesis of a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M includes at least one of 3d transition metals, and a sintering step of sintering and reacting the precursor obtained by the mixing step, wherein an electrically conductive agent is added to the starting material for synthesis or to the precursor.

In the method for producing a positive electrode active material according to the present invention, in which an electrically conductive agent is added to the starting material for synthesis or to the precursor, the produced positive electrode active material exhibits superior load characteristics and electrode molding performance, to realize a high capacity.

The present invention also provides a method for producing a non-aqueous electrolyte secondary battery having a positive electrode active material capable of reversibly doping/undoping lithium, a negative electrode mounted facing the positive electrode and capable of reversibly doping/undoping lithium, and a non-aqueous electrolyte interposed between the positive electrode and the negative electrode, wherein the positive electrode active material is synthesized by a mixing step of

mixing a plurality of substances, proving a starting material for synthesis of a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M includes at least one of 3d transition metals, to give a precursor, and a sintering step of sintering and reacting the precursor obtained by the mixing step, and wherein an electrically conductive agent is added to the starting material for synthesis or to the precursor to synthesize the positive electrode active material.

In the method for producing the non-aqueous electrolyte secondary battery according to the present invention, an electrically conductive agent is added to the starting material for synthesis or to the precursor. Since the positive electrode active material exhibits satisfactory load characteristics and electrode molding performance, the produced non-aqueous electrolyte secondary battery is of a high capacity.

Brief Description of the Drawings

Fig.1 is a cross-sectional view showing an illustrative structure of a non-aqueous electrolyte secondary battery embodying the present invention.

Fig.2 is a graph showing a powder X-ray diffraction pattern of LiFePO_4 synthesized in Example 1.

Fig.3 is a graph showing a powder X-ray diffraction pattern of LiFePO_4 synthesized in Comparative Example 1.

Fig.4 is a graph showing charging/discharging characteristics of the battery of

Example 1.

Fig.5 is a graph showing the relation between the number of cycles and the charging/discharging capacity of the battery of Example 1.

Fig.6 is a graph showing the charging/discharging characteristics of the battery of Comparative Example 1.

Fig.7 is a graph showing the relation between the number of cycles and the charging/discharging capacity of the battery of Comparative Example 1.

Fig.8 is a graph showing a powder X-ray diffraction pattern of LiFePO_4 synthesized in Example 3.

Fig.9 is a graph showing a powder X-ray diffraction pattern of LiFePO_4 synthesized in Comparative Example 2.

Fig.10 is a graph showing the charging/discharging characteristics of the battery of Example 3.

Fig.11 is a graph showing the relation between the number of cycles and the charging/discharging capacity of the battery of Example 3.

Fig.12 is a graph showing the relation between the sintering temperature in synthesizing LiFePO_4 .

Fig.13 is a graph showing the charging/discharging characteristics of the battery of Comparative Example 2.

Fig.14 is a graph showing the relation between the number of cycles and the charging/discharging capacity of the battery of Comparative Example 2.

Fig.15 is a graph showing a powder X-ray diffraction pattern of an LiFePO_4 carbon complex material synthesized in Example 7.

Fig.16 is a graph showing a powder X-ray diffraction pattern of an LiFePO_4 carbon complex material synthesized in Example 8.

Fig.17 is a graph showing a powder X-ray diffraction pattern of an $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$ carbon complex material synthesized in Example 9.

Fig.18 is a graph showing a powder X-ray diffraction pattern of an LiFePO_4 carbon complex material synthesized in Comparative Example 7.

Fig.19 is a graph showing the charging/discharging characteristics of the battery of Example 7.

Fig.20 is a graph showing the relation between the number of cycles and the charging/discharging capacity of the battery of Example 7.

Fig.21 is a graph showing load characteristics of a battery of Example 8.

Fig.22 is a graph showing the charging/discharging characteristics of the battery of Example 9.

Fig.23 is a graph showing load characteristics of batteries of Example 7 and Comparative Example 1.

Fig.24 is a graph showing the charging/discharging characteristics of the battery of Comparative Example 4.

Fig.25 is a graph showing powder X-ray diffraction patterns of LiFePO_4 carbon complex material synthesized in Examples 10 to 12.

Fig.26 is a graph showing powder an X-ray diffraction pattern of LiFePO₄ synthesized in Comparative Example 5.

Fig.27 is a graph showing the charging/discharging characteristics of the battery of Example 11.

Fig.28 is a graph showing the charging/discharging characteristics of the battery of Example 12.

Fig.29 is a graph showing the charging/discharging characteristics of the battery of Comparative Example 4.

Fig.30 is a graph showing volumetric grain size distribution of the LiMnPO₄ carbon complex material of Examples 10 to 12 and LiMnPO₄ of Comparative Example 5.

Fig.31 is a graph showing the volumetric grain size distribution of the LiMnPO₄ carbon complex material of Example 12 in integrated values of the passed portions.

Best mode for Carrying out the Invention

The present invention is hereinafter explained in detail.

A non-aqueous electrolyte secondary battery, manufactured in accordance with the present invention, is a so-called lithium secondary battery, and includes, as basic constituent elements, a positive electrode, containing a positive electrode active material, a negative electrode and a non-aqueous electrolyte.

This positive electrode active material contains a compound having an olivinic

structure and which is represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$, where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M contains at least one of 3d transition metals.

The compounds represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ may be enumerated by, for example, $\text{Li}_x\text{Fe}_y\text{PO}_4$, $\text{Li}_x\text{Mn}_y\text{PO}_4$, $\text{Li}_x\text{Co}_y\text{PO}_4$, $\text{Li}_x\text{Ni}_y\text{PO}_4$, $\text{Li}_x\text{Cu}_y\text{PO}_4$, $\text{Li}_x(\text{Fe, Mn})_y\text{PO}_4$, $\text{Li}_x(\text{Fe, Co})_y\text{PO}_4$, $\text{Li}_x(\text{Fe, Ni})_y\text{PO}_4$, $\text{Li}_x(\text{Cu, Mn})_y\text{PO}_4$, $\text{Li}_x(\text{Cu, Co})_y\text{PO}_4$ and $\text{Li}_x(\text{Cu, Ni})_y\text{PO}_4$, where the proportions of elements in parentheses () are arbitrary.

The compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ preferably contains particles with the particle size not larger than $10 \mu\text{m}$. On the other hand, the specific surface area of the compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ is preferably not less than $0.5 \text{ m}^2/\text{g}$.

In preparing a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ as the aforementioned positive electrode active material, plural substances as starting material for synthesis of the compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ are mixed together to prepare a precursor by a mixing process. The precursor prepared in the mixing process is then sintered and reacted by a sintering process. A reducing agent is added to the precursor in the mixing process.

The method for preparing LiFePO_4 , as the compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$, is now explained.

In the mixing process, iron oxalate (FeC_2O_4), as a starting material for synthesis,

ammonium hydrogen phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$) and lithium carbonate (Li_2CO_3) are mixed together in a pre-set ratio and added to with e.g., iron powders (Fe) as a reducing agent. These substances are mixed sufficiently to give a precursor.

In the sintering process, this precursor is sintered at a pre-set temperature in an atmosphere of an inert gas, such as nitrogen. This synthesizes LiFePO_4 .

If air is left in the precursor in synthesizing LiFePO_4 , Fe^{2+} in iron oxalate, as a bivalent iron compound, tends to be oxidized with oxygen in air to Fe^{3+} . As a result, the trivalent iron compound tends to be mixed as an impurity into the synthesized LiFePO_4 .

According to the present invention, iron powders are added as a reducing agent in the mixing process. If Fe^{2+} in iron oxalate, as a bivalent iron compound, is oxidized to Fe^{3+} by oxygen in air contained in the precursor, the iron powders contained in the precursor reduces this Fe^{3+} to Fe^{2+} . Since this prohibits the trivalent iron compound from mixing into the synthesized LiFePO_4 , it is possible to produce single-phase LiFePO_4 .

The iron powder, added as a reducing agent, forms a part of the starting material for synthesis, and is synthesized into LiFePO_4 as an ultimate product. If the reducing agent becomes a part of the starting material for synthesis of LiFePO_4 , there is no fear of mixing of impurities, so that single-phase LiFePO_4 may be produced. If the reducing agent is used as a part of LiFePO_4 , the single-phase LiFePO_4 may be produced without the risk of mixing of impurities. Moreover, if the reducing agent is

used as a part of LiFePO₄, there is no necessity of removing the reducing agent left over after the end of the reaction to render it possible to synthesize LiFePO₄ efficiently.

If iron powders are used as a part of the reducing agent or of the starting material for LiFePO₄ synthesis, these iron powders are desirably added to the precursor in an amount ranging from 1 wt% to 30 wt% based on the sum total of Fe in the iron starting material. If the amount of addition of the iron powders is less than this value, oxidation of Fe²⁺ cannot be prevented sufficiently. On the other hand, since the iron powders (Fe) are lower in reactivity than Fe²⁺ in iron oxalate (FeC₂O), the reaction of synthesis of LiFePO₄ does not proceed sufficiently if the amount of the iron powders exceeds 30 wt%. By setting the amount of addition of the iron powders in a range from 1 to 30 wt% based on the weight of Fe in the iron starting material, it is possible to prohibit oxidation of Fe²⁺ to produce the single-phase LiFePO₄ without obstructing the synthetic reaction.

As the reducing agent, oxalic acid, formic acid or hydrogen may be used in place of iron powders, insofar as no solid product is left over after the sintering process.

In the present technique, in which the reducing agent is added to the precursor as described above to suppress generation of trivalent iron compounds by trace oxygen, and the precursor is sintered in a nitrogen stream, it has become possible to synthesize LiFePO₄ at a temperature e.g., of 300 °C which is appreciably lower than

the conventional temperature of 800°C. Stated differently, LiFePO₄ can be synthesized over a broader temperature range than heretofore to increase the latitude of selection of the precursor sintering temperature, referred to below simply sintering temperature. If the sintering temperature is as high as 800°C, as in the conventional system, the energy consumption is correspondingly increased, whilst the load imposed on reaction apparatus etc is also increased.

The present inventors have directed attention to the relation between the precursor sintering temperature in synthesizing LiFePO₄ and the battery employing LiFePO₄ as the active material and investigated into an optimum sintering temperature for synthesizing LiFePO₄ in realization of a high capacity.

As a result, it has been found that the sintering temperature for sintering the LiFePO₄ precursor is desirably not lower than 350°C and not higher than 790°C. If the sintering temperature is lower than 350°C, there is a risk that the chemical reaction and crystallization do not proceed sufficiently, such that homogeneous LiFePO₄ cannot be produced. On the other hand, if the sintering temperature is higher than 790°C, there is a risk that crystallization proceeds excessively to obstruct lithium diffusion. So, by sintering the precursor at a temperature not lower than 350°C and not higher than 790°C, to synthesize LiFePO₄, homogeneous LiFePO₄ can be produced to realize a high capacity exceeding 120 mAh/g which is the capacity of LiFePO₄ produced by the conventional method.

It is more desirable that the sintering temperature be not lower than 450°C and

not higher than 700 °C. By sintering the precursor in a temperature range from 450 °C to 700 °C, it is possible to realize the high real capacity approaching to 170 mAh/g which is the theoretical capacity of LiFePO_4 .

In the above-described manufacturing method for the positive electrode active material, in which a reducing agent is added to the precursor in synthesizing $\text{Li}_x\text{M}_y\text{PO}_4$, it is possible to prevent oxidation of M as a 3d transition metal, so that the single-phase $\text{Li}_x\text{M}_y\text{PO}_4$, can be produced without mixing of impurities. Moreover, $\text{Li}_x\text{M}_y\text{PO}_4$, can be produced at a sintering temperature lower than in the conventional system. So, with the present manufacturing method for the positive electrode active material, $\text{Li}_x\text{M}_y\text{PO}_4$ can be produced with which it is possible to realize a high capacity.

By using $\text{Li}_x\text{M}_y\text{PO}_4$, synthesized as described above, it is possible to produce a non-aqueous electrolyte secondary battery of high capacity and superior cyclic characteristics and lithium ion doping/undoping performance.

On the other hand, $\text{Li}_x\text{M}_y\text{PO}_4$ as a positive electrode active material can be produced as follows: First, plural materials as a starting material for synthesis of a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ are mixed to give a precursor by a mixing process. Then, air contained in the precursor obtained by the mixing process is removed by way of a de-aerating process. The precursor freed of air by the de-aerating process is sintered and reacted by a sintering process to produce $\text{Li}_x\text{M}_y\text{PO}_4$.

Another method for synthesizing e.g., LiFePO_4 as a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ is hereinafter explained.

In the mixing process, iron acetate ($\text{Fe}(\text{CH}_3\text{COO})_2$), hydrogen ammonium phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$) and lithium carbonate (Li_2CO_3) are mixed at a pre-set ratio to give the precursor.

In the de-aerating process, the precursor is homogenized sufficiently and de-aerated to remove air contained in the precursor. As the de-aerating processing, the atmosphere of the precursor is replaced by vacuum to introduce an inert gas. This removes air contained in the precursor. As another example of the de-aerating process, a solvent with a boiling point not higher than 250°C is made to co-exist with the precursor to vaporize off the solvent into the inert gas. This removes air contained in the precursor. The solvent boiling at 250°C or lower may be exemplified by, for example, water and ethanol.

In the sintering process, the precursor from the de-aerating processing is sintered at a pre-set temperature in an atmosphere of an inert gas, such as nitrogen. This synthesizes LiFePO_4 .

If air is left in the precursor at the time of synthesis of LiFePO_4 , Fe^{2+} in iron acetate as a bivalent iron compound may be oxidized by oxygen in air and thereby turned into Fe^{3+} . So, the trivalent iron compound may be mixed as an impurity in the synthesized LiFePO_4 .

According to the present invention, the air contained in the precursor is removed by the de-aerating process to prevent oxidation of Fe^{2+} contained in iron acetate. In this manner, no trivalent iron compound is mixed into synthesized LiFePO_4 .

to make it possible to produce single-phase LiFePO₄.

The sintering temperature, which is the temperature of sintering the precursor in synthesizing LiFePO₄, is desirably not lower than 350°C and not higher than 790°C, as in the case of the range of the sintering temperature.

In the above-described manufacturing method for the positive electrode active material, the air contained in the precursor is removed in synthesizing Li_xM_yPO₄, thus preventing oxidation of M as a 3d transition metal. This makes it possible to produce single-phase LiFePO₄ free of impurities. Moreover, Li_xM_yPO₄ can be synthesized at a lower sintering temperature. Thus, with the present manufacturing method for the positive electrode active material, Li_xM_yPO₄, realizing a high capacity, may be produced.

By employing Li_xM_yPO₄, synthesized as described above, as the positive electrode active material, it is possible to produce a non-aqueous electrolyte secondary battery of high capacity and superior cyclic characteristics and lithium ion doping/undoping performance.

Moreover, in synthesizing a compound sample composed of Li_xM_yPO₄ and an electrically conductive material, as a positive electrode active material containing Li_xM_yPO₄, plural materials as a starting material for synthesis of Li_xM_yPO₄ are mixed to give a precursor by a mixing process. The precursor obtained by the mixing process is then sintered and reacted by way of a sintering process. At this time, an electrically conductive material is added to the starting material for synthesis or to the precursor.

This electrically conductive material may be exemplified by carbon, silver or an electrically conductive high polymer material. The carbon black may be exemplified by, for example, graphite, acetylene black or the like carbon black material.

The electrically conductive material is desirably added in a range of 0.5 to 20 parts by weight to 100 parts by weight of $\text{Li}_x\text{M}_y\text{PO}_4$. If the amount of the electrically conductive material is less than 0.5 wt%, it is likely that no sufficient effect can be achieved. If conversely the amount of the electrically conductive material exceeds 20 wt%, the proportion of M as the main partner of the oxidation in the positive electrode active material is low such that the energy density of the non-aqueous electrolyte secondary battery is likely to be lowered.

Therefore, if, in the positive electrode active material, the electrically conductive material is added in a range of 0.5 to 20 parts by weight to 100 parts by weight of $\text{Li}_x\text{M}_y\text{PO}_4$, load characteristics and electrode molding performance are improved, such that the non-aqueous electrolyte secondary battery having this compound sample as the positive electrode active material has a high capacity.

As a manufacturing method for a compound sample as the positive electrode active material, a manufacturing method for synthesizing the LiFePO_4 carbon compound material composed of LiFePO_4 and carbon as the $\text{Li}_x\text{M}_y\text{PO}_4$ and as the electrically conductive material is hereinafter explained.

In adding carbon to the precursor of the LiFePO_4 , iron oxalate (FeC_2O_4),

ammonium hydrogen phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$) and lithium carbonate (Li_2CO_3) are sufficiently mixed in the mixing process to give a precursor, which precursor then is calcined at a lower temperature in an atmosphere of an inert gas, such as nitrogen. The calcined precursor and the carbon are mixed together and pulverized. In the sintering process, sintering is made at a pre-set temperature in an atmosphere of an inert gas, such as nitrogen, to produce a LiFePO_4 carbon compound material.

If carbon is added to the starting material for synthesis of LiFePO_4 , carbon is added and mixed at the outset in the mixing process to a starting material for synthesis composed of iron oxalate (FeC_2O_4), ammonium hydrogen phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$) and lithium carbonate (Li_2CO_3) and the resulting mixture is calcined at a lower temperature in an atmosphere of an inert gas, such as nitrogen. In the sintering process, the calcined mixture is sintered at a pre-set temperature in the inert gas atmosphere, such as nitrogen, to produce an LiFePO_4 carbon compound material.

In the present technique, the starting material for synthesis is mixed thoroughly as described above to prepare the precursor. By thoroughly mixing the starting material for synthesis, the respective components of the starting material are mixed evenly to provide a precursor having an increased number of contact points. This precursor is sintered and synthesized in a nitrogen stream. This makes it possible to synthesize the compound sample by sintering the precursor at a temperature of e.g., 300 °C which is appreciably lower than 800 °C which is the sintering temperature for synthesizing LiFePO_4 by the conventional method.

Stated differently, the LiFePO₄ carbon compound material can be synthesized over a wide temperature range to provide for a wider temperature selection range for synthesis. This sintering temperature is preferably not less than 350 °c and not higher than 790 °C and more preferably not less than 450 °c and not higher than 700 °C.

If a compound sample composed of a compound of Li_xM_yPO₄, where M is Mn the redox potential of which is higher than that of Fe, and an electrically conductive material, for example, a compound material composed of LiFe_xMn_{1-x}PO₄ and carbon, is to be synthesized as the positive electrode active material, iron oxalate (FeC₂O₄), ammonium hydrogen phosphate (NH₄H₂PO₄) and lithium carbonate (Li₂CO₃), and manganese acetate tetrahydride (Mn(CH₃OO)₂·4H₂O) or manganese carbonate (MnCO₃) are mixed at a pre-set ratio. Otherwise, the technique of synthesizing the LiFePO₄ carbon compound material is used to prepare LiFe_xMn_{1-x}PO₄.

If an Li_xM_yPO₄ carbon compound material, composed of Li_xM_yPO₄ and carbon where x is such that 0 < x ≤ 2 and y is such that 0.8 ≤ y ≤ 1.2, is to be synthesized as the positive electrode active material, the above-described technique of synthesizing the LiFePO₄ carbon compound material is followed except that manganese carbonate (MnCO₃), ammonium hydrogen phosphate (NH₄H₂PO₄) and lithium carbonate (Li₂CO₃) are mixed thoroughly at a pre-set ratio for use as a precursor, to prepare the Li_xM_yPO₄ precursor.

The conventional olivinic Li_xM_yPO₄ is low in electrically conductivity to render oxidation of Mn difficult. So, the present inventors added carbon to the starting

material for synthesis of the olivinic $\text{Li}_x\text{Mn}_y\text{PO}_4$ or its precursor to synthesize the $\text{Li}_x\text{Mn}_y\text{PO}_4$ carbon compound material. Since oxidation of Mn newly occurs in this $\text{Li}_x\text{Mn}_y\text{PO}_4$ carbon compound material, it is possible to use this material as the positive electrode active material.

In the above-described manufacturing method for the positive electrode active material, in which the compound sample composed of $\text{Li}_x\text{Mn}_y\text{PO}_4$ and the electrically conductive material is to be synthesized, it is possible to synthesize a positive electrode active material having load characteristics and electrode molding performance more desirable than those of the positive electrode active material composed solely of LiM_yPO_4 . Moreover, with the positive electrode active material, thus prepared, superior load characteristics and electrode molding performance are achieved, even if the amount of the electrically conductive material newly added at the time of preparing the electrodes is small, thus enabling the use of an electrically conductive agent having a large volumetric density, such as graphite.

So, the non-aqueous electrolyte secondary battery, containing this compound sample as the positive electrode active material, is of high capacity and superior cyclic characteristics, because migration of electrons occurs smoothly in the electrodes. On the other hand, there is no necessity of adding a new electrically conductive agent to the positive electrode mixture of the non-aqueous electrolyte secondary battery, thus increasing the energy density.

If, in the non-aqueous electrolyte secondary battery employing the positive

electrode active material, synthesized as described above, the electrolyte used in a non-aqueous electrolyte secondary battery 1 is liquid, the non-aqueous electrolyte secondary battery 1 includes a positive electrode 2, a positive electrode can 3, accommodating the positive electrode 2, a negative electrode 4, a negative electrode can 5, accommodating the negative electrode 4, a separator 6 mounted between the positive electrode 2 and the negative electrode 4 and an insulating gasket 7. The non-aqueous electrolyte is charged in the positive electrode 2 and in the negative electrode can 5.

The positive electrode 2 is comprised of a positive electrode collector of, for example, an aluminum foil, and a layer of a positive electrode active material containing the positive electrode active material, synthesized as described above. As a binder contained in the layer of the positive electrode active material, it is possible to use a known resin material routinely used as a binder for the layer of the positive electrode active material for this type of the non-aqueous electrolyte secondary battery.

The positive electrode can 3, accommodating the positive electrode 2, serves as an external positive electrode for the non-aqueous electrolyte secondary battery 1.

The negative electrode 4 is formed e.g., by a foil of metal lithium operating as a negative electrode active material. If a material capable of doping/undoping lithium is used as the negative electrode active material, the negative electrode 4 is the layer of the negative electrode active material containing the negative electrode active material and the negative electrode collector on which is formed the layer of the

negative electrode active material. As the negative electrode collector, a nickel foil, for example, is used. As the negative electrode active material capable of doping/undoping lithium, metal lithium, lithium alloys, an electrically conductive high polymer material doped with lithium, and a laminated compound, such as a carbon material or a metal oxide, may be used. As a binder contained in the layer of the negative electrode active material, any suitable known binders routinely used as a binder for the layer of the negative electrode active material for this sort of the non-aqueous electrolyte secondary battery may be used.

The negative electrode can 5, accommodating the negative electrode 4, serves as an external positive electrode for the non-aqueous electrolyte secondary battery 1.

The separator 6, used for separating the positive electrode and the negative electrode from each other, may be formed of any suitable known materials routinely used as a separator for the layer of the negative electrode active material for this sort of the non-aqueous electrolyte secondary battery. For example, a film of a high molecular material, such as polypropylene, is used. From the relation between conductivity of lithium ions and the energy density, it is necessary that the thickness of the separator 6 be as thin as possible. Specifically, the thickness of the separator 6 of, for example, not more than 50 μm , is desirable.

The insulating gasket 7 is built into and unified with the negative electrode can 5 and serves for preventing leakage of the non-aqueous electrolytic solution charged into the positive electrode can 3 and the negative electrode can 5.

As the non-aqueous electrolytic solution, a solution obtained on dissolving the electrolyte in a non-protonic non-aqueous solvent is used.

The non-aqueous solvent may be exemplified by, for example, propylene carbonate, ethylene carbonate, butylene carbonate, vinylene carbonate, γ -butyrolactone, sulforane, 1,2-dimethoxyethane, 1,2-diethoxyethane, 2-methyltetrahydrofuran, 3- methyl 1,3- dioxorane, methyl propionate, methyl lactate, dimethyl carbonate, diethyl carbonate and dipropyl carbonate. Especially, from voltage stability, cyclic carbonates, such as propylene carbonate or vinylene carbonate, or chain carbonates, such as dimethyl carbonate, diethyl carbonate or dipropyl carbonate, are preferably used. As this non-aqueous solvent, only one type non-aqueous solvent or a mixture of two or more non-aqueous solvents may be used.

As the electrolyte, dissolved in the non-aqueous solvent, lithium salts, such as LiPF₆, LiClO₄, LiAsF₆, LiBF₄, LiCF₃SO₃ or LiN(CF₃SO₂)₂, may be used. Of these lithium salts, LiPF₆ or LiBF₄ may preferably be used.

The non-aqueous electrolyte secondary battery 1 is manufactured, e.g., by the following method:

For preparing the positive electrode 2, the positive electrode active material and the binder are dispersed in a solvent to prepare a slurred positive electrode mixture. The so-prepared positive electrode mixture then is evenly coated on a current collector and dried in situ to prepare a layer of the positive electrode active material to complete the positive electrode 2. For preparing the negative electrode 4, the negative electrode

active material and the binder are dispersed in a solvent to prepare a slurried negative electrode mixture. The so-prepared negative electrode mixture then is evenly coated on a current collector and dried in situ to prepare a layer of the negative electrode active material to complete the negative electrode 4. The non-aqueous electrolytic solution is prepared by dissolving an electrolyte salt in a non-aqueous solvent.

The positive electrode 2 and the negative electrode 4 are accommodated in the positive electrode can 3 and in the negative electrode can 5. The separator 6 formed e.g., by a porous polypropylene film is arranged between the positive electrode 2 and the negative electrode 4. The non-aqueous electrolytic solution is charged into the positive electrode can 3 and the negative electrode can 5. The electrode cans 3, 5 are caulked fixedly through the insulating gasket 7 to complete the non-aqueous electrolyte secondary battery 1.

In the above-described embodiment, the non-aqueous electrolyte secondary battery 1 employing the non-aqueous electrolytic solution is used, as an example, as the non-aqueous electrolyte secondary battery. The present invention is, however, not limited to this and may be applied to the use as the non-aqueous electrolyte of a solid electrolyte or a gelated solid electrolyte containing a swelling solvent. The present invention may also be applied to a variety of shapes of the non-aqueous electrolyte secondary batteries, such as a cylindrical shape, a square shape, a coin or a button shape, or to a variety of sizes of the non-aqueous electrolyte secondary battery, such as a thin type or large-sized batteries.

In the above-described embodiment, the manufacturing method for an positive electrode active material by synthesizing LiFePO_4 , an LiFePO_4 carbon compound, a compound sample composed of $\text{LiFe}_{x}\text{Mn}_{1-x}\text{PO}_4$ and carbon or a $\text{Li}_x\text{Mn}_y\text{PO}_4$ carbon compound material composed of $\text{Li}_x\text{Mn}_y\text{PO}_4$ and carbon, has been explained. However, it is only necessary that the positive electrode active material be based on $\text{Li}_x\text{M}_y\text{PO}_4$ as the basic composition, such that any element may be added to or used as a substituent in $\text{Li}_x\text{M}_y\text{PO}_4$ as long as the reaction and crystallization proceed at a sintering temperature not lower than 350°C and not higher than 790°C , as described above. $\text{Li}_x\text{M}_y\text{PO}_4$ may also suffer from defects.

Also, in the above-described embodying the present invention of the present invention, a solid-phase reaction of mixing and sintering a powdered material as a starting material for synthesis of the compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ in the preparation of the positive electrode active material is explained as an example. The present invention, however, is not limited to this and may be applied to the use of various other methods for chemical synthesis than the solid-phase reaction.

Meanwhile, a starting material of high reactivity is preferably used as the starting material for synthesis of $\text{Li}_x\text{M}_y\text{PO}_4$. For example, as the starting material for synthesis of LiFePO_4 , various starting materials, such as lithium hydroxide, lithium nitrate, lithium acetate, lithium phosphate, iron (II) phosphate or iron (II) oxide may be used in addition to the above-mentioned compounds.

In the following, specified Examples and Comparative Examples, embodying the present invention, are explained based on the experimental results.

<Experiment 1>

In an experiment 1, in synthesizing LiFePO₄ as the positive electrode active material, an Example in which iron powders were added as a reducing agent to the precursor and a Comparative Example in which iron powders were not added to the precursor, were synthesized. Plural non-aqueous electrolyte secondary batteries, employing these positive electrode active materials, were prepared, and evaluation was made of difference in characteristics of the respective positive electrode active materials and the non-aqueous electrolyte secondary batteries.

Example 1

First, LiFePO₄ was synthesized. For synthesizing LiFePO₄, ammonium dihydrogen phosphate (NH₄H₂PO₄) as a starting material of a coarser crystallite size was sufficiently pulverized at the outset. Then, iron powders (Fe), iron oxalate dihydride (FeC₂O₄·2HO), ammonium dihydrogen phosphate (NH₄H₂PO₄) and lithium carbonate (Li₂CO₃) were mixed together to a molar ratio of 0.4: 1.6: 2: 1 for thirty minutes in a mortar to give a precursor. This precursor then was calcined in a nitrogen atmosphere at 300°C for 12 hours. The so-calcined precursor was sintered in the nitrogen atmosphere at 600°C for 24 hours to synthesize LiFePO₄.

The so-prepared LiFePO₄ was used as the positive electrode active material to prepare a battery. 70 wt% of dried LiFePO₄, as the positive electrode active material,

25 wt% of acetylene black, as an electrically conductive material, and 5 wt% of polyvinylidene fluoride, as a binder, were evenly mixed into dimethyl formamide as a solvent to prepare a paste-like positive electrode mixture. Meanwhile, #1300 manufactured by Aldrich Inc. was used as the polyvinylidene fluoride. This positive electrode mixture was applied to an aluminum mesh, as a current collector, and dried in situ in a dry argon atmosphere at 100 °C for one hour to form a layer of the positive electrode active material. The aluminum mesh, on which the layer of the positive electrode active material was formed, was punched to a disc 15.5 mm in diameter to form a pellet-like positive electrode. Meanwhile, this positive electrode carries 60 mg of the active material.

A metal lithium foil was punched to substantially the same shape as the positive electrode and used as a negative electrode. In a mixed solvent of equal parts in volume of propylene carbonate and dimethyl carbonate was dissolved LiPF₆ at a concentration of 1 mol/l to prepare a non-aqueous electrolytic solution.

The positive electrode, prepared as described above, was accommodated in the positive electrode can, whilst the negative electrode was accommodated in the negative electrode can and the separator was arranged between the positive electrode and the negative electrode. The non-aqueous electrolytic solution was charged into the positive electrode can and the negative electrode can. The electrode cans 3, 5 are caulked fixedly through the insulating gasket 7 to complete a 2025 type coin-shaped test cell.

Example 2

LiFePO_4 was prepared in the same way as in Example 1, except using iron acetate in place of iron oxalate as the starting material for synthesis and mixing iron powders (Fe), iron acetate ($\text{FeCH}_3\text{COO}_2$), ammonium dihydrogen phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$) and lithium carbonate (Li_2CO_3) to a molar ratio of 0.2: 1.8: 2: 1. The so-produced LiFePO_4 was used as the positive electrode active material to prepare a test cell.

Comparative Example 1

LiFePO_4 was prepared in the same way as in Example 1, except adding no reducing agent and mixing iron oxalate dihydride ($\text{FeC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$), ammonium dihydrogen phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$) and lithium carbonate (Li_2CO_3) to a molar ratio of 2: 2: 1. The so-produced LiFePO_4 was used as the positive electrode active material to prepare a test cell.

Then, measurement was made of the powder X-ray diffraction pattern of the LiFePO_4 prepared by the above-described method. The measurement conditions of the powder X-ray diffraction were as follows:

apparatus used: RIGAKU RINT 2500 rotary counter pair negative electrode

goniometer: vertical type standard, radius 185 mm

counter monochromator: used

filter: not used

slit width

divergent slit (DS) = 1°

receiving slit (RS) = 1°

scattering slit (SS) = 0.15 mm

counter device: scintillation counter

measurement method: reflection method, continuous scan

scanning range: $2\theta = 10^\circ$ to 80°

scanning speed: 4°/minute

The powder X-ray diffraction pattern of LiFePO_4 , synthesized in Example 1, is shown in Fig.2, from which it is seen that a single-phase LiFePO_4 has been obtained since the presence of the impurity other than LiFePO_4 is not confirmed in the product.

The precursor was sintered at plural temperature points in a range from 320°C to 850°C to prepare LiFePO_4 in a similar manner. It was found that there was no impurity in LiFePO_4 synthesized in the range from 320°C to 850°C so that a single-phase LiFePO_4 has been obtained.

On the other hand, the presence of the impurity other than LiFePO_4 was not noticed in the product synthesized in Example 2 from the powder X-ray diffraction pattern of LiFePO_4 synthesized in Example 2, thus testifying to the generation of the single-phase LiFePO_4 .

The powder X-ray diffraction pattern of LiFePO_4 , synthesized in Comparative Example 2, is shown in Fig.3, from which it is seen that impurities other than LiFePO_4 are present in the product such that no single-phase LiFePO_4 has been yielded.

Meanwhile, in Fig.3, a peak attributable to the impurity is marked with an asterisk *. This impurity is presumably a trivalent iron compound, such as Fe_2O_3 .

The test cells, prepared as described above, were subjected to the charging/discharging test, in which each test cell was charged by constant current charging and, when the battery voltage reached 4.5V, the charging system was switched from the constant current charging to constant voltage charging, and charging was carried out as the voltage of 4.5 V was kept. The charging was stopped when the current fell below 0.01 mA/cm². The discharging then was carried out and stopped at a time point when the battery voltage was lowered to 2.0 V. Meanwhile, charging/discharging was carried out at ambient temperature (23 °C), with the current density at this time being 0.12 mA/cm².

The charging/discharging characteristics of the battery of Example 1 are shown in Fig.4, from which it is seen that the battery of Example 1 showed a flat potential in the vicinity of 3.4 V, thus indicating that the high reversible charging/discharging capacity of 163 mAh/g, which is close to the theoretical capacity of 170 mAh/g, is being produced.

The relation between the number of cycles and the charging/discharging capacity of the battery of Example 1 is shown in Fig.5, from which it is seen that the battery of the Example 1 undergoes cyclic deterioration as low as 0.1%/cycle, thus demonstrating stable battery characteristics.

The charging/discharging characteristics of the battery of Comparative Example

1 are shown in Fig.6. The relation between the number of cycles and the charging/discharging capacity of the battery of Comparative Example 1 is shown in Fig.7. From Figs.6 and 7, it may be seen that the battery of Comparative Example 1 is satisfactory in cyclic characteristics, however, the battery of Comparative Example 1 suffers from significant capacity loss at an initial stage, with the charging/discharging efficiency being low.

Thus, it may be seen that, by adding a reducing agent to the precursor in synthesizing LiFePO_4 , Fe^{2+} in the precursor may be prevented from being oxidized by the residual oxygen to become Fe^{3+} to give a single-phase LiFePO_4 . The battery employing this single-phase LiFePO_4 as the positive electrode active material is superior in charging/discharging characteristics and in cyclic characteristics.

<Experiment 2>

In this experiment 2, in synthesizing LiFePO_4 as the positive electrode active material, an Example in which a precursor was de-aerated, and a Comparative Example in which iron powders were not added, were synthesized, and plural non-aqueous electrolyte secondary batteries employing these positive electrode active materials were prepared. Of the respective positive electrode active materials and the non-aqueous electrolyte secondary batteries, the difference in characteristics was evaluated.

Example 3

First, LiFePO_4 was synthesized. For synthesizing LiFePO_4 , ammonium

dihydrogen phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$) as a starting material of a coarser crystallite size was sufficiently pulverized at the outset. Then, iron acetate ($\text{FeCH}_3\text{COO}_2$), ammonium dihydrogen phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$) and lithium carbonate (Li_2CO_3) were mixed together to a molar ratio of 2: 2: 1 for thirty minutes in a mortar to give a precursor.

This precursor then was charged into an electrical oven and, after the atmosphere was replaced by vacuum, a nitrogen gas was introduced into the oven. This precursor was calcined in a nitrogen atmosphere at 300°C for 12 hours, after which the precursor was sintered in the nitrogen atmosphere at 600°C for 24 hours to synthesize LiFePO_4 .

The so-prepared LiFePO_4 was used as a positive electrode active material to produce a test cell in the same way as in Example 1.

Example 4

First, LiFePO_4 was synthesized. For synthesizing LiFePO_4 , ammonium dihydrogen phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$) as a starting material of a coarser crystallite size was sufficiently pulverized at the outset. Then, iron acetate ($\text{FeCH}_3\text{COO}_2$), ammonium dihydrogen phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$) and lithium carbonate (Li_2CO_3) were mixed together to a molar ratio of 2: 2: 1 for thirty minutes in a mortar to give a precursor. The resulting mixture was further mixed for 48 hours, using zirconia balls 2mm in diameter, with water as a solvent.

This mixture then was dried to a clayey state, after which the mixture was

sintered in a nitrogen stream in an electrical oven at 120 °C to remove residual water to produce a precursor with high homogeneity. This precursor then was calcined at 300 °C for 12 hours in a nitrogen atmosphere and further sintered in the nitrogen atmosphere at 600 °C for 24 hours to synthesize LiFePO₄.

Using the so-prepared LiFePO₄ as the positive electrode active material, a test cell was prepared as in Example 1.

Example 5

First, LiFePO₄ was synthesized. For synthesizing LiFePO₄, ammonium dihydrogen phosphate (NH₄H₂PO₄) as a starting material of a coarser crystallite size was sufficiently pulverized at the outset. Then, iron acetate (FeCH₃COO)₂, ammonium dihydrogen phosphate (NH₄H₂PO₄) and lithium carbonate (Li₂CO₃) were mixed together to a molar ratio of 2: 2: 1 for thirty minutes in a mortar to give a precursor. The resulting mixture was further mixed for 48 hours, using zirconia balls 2 mm in diameter, with ethanol as a solvent.

This mixture then was dried to a clayey state, after which the mixture was sintered in a nitrogen stream in an electrical oven at 120 °C to remove residual ethanol to produce a precursor with extremely high homogeneity. This precursor then was calcined at 300 °C for 12 hours in a nitrogen atmosphere and further sintered in the nitrogen atmosphere at 600 °C for 24 hours to synthesize LiFePO₄.

Using the so-prepared LiFePO₄ as the positive electrode active material, a test cell was prepared as in Example 1.

Example 6

First, LiFePO₄ was synthesized. For synthesizing LiFePO₄, ammonium dihydrogen phosphate (NH₄H₂PO₄) as a starting material of a coarser crystallite size was sufficiently pulverized at the outset. Then, iron acetate (FeCH₃COO)₂, ammonium dihydrogen phosphate (NH₄H₂PO₄) and lithium carbonate (Li₂CO₃) were mixed together to a molar ratio of 2: 2: 1 for thirty minutes in a mortar to give a precursor. The resulting mixture was further mixed for 48 hours, using zirconia balls 2 mm in diameter, with acetone as a solvent.

This mixture then was dried to a clayey state, after which the mixture was sintered in a nitrogen stream in an electrical oven at 120 °C to remove residual acetone to produce a precursor with high homogeneity. This precursor then was calcined at 300 °C for 12 hours in a nitrogen atmosphere and further sintered in the nitrogen atmosphere at 600 °C for 24 hours to synthesize LiFePO₄.

Using the so-prepared LiFePO₄ as the positive electrode active material, a test cell was prepared as in Example 1.

Comparative Example 2

First, LiFePO₄ was synthesized. For synthesizing LiFePO₄, iron acetate (FeCH₃COO)₂, ammonium dihydrogen phosphate (NH₄H₂PO₄) and lithium carbonate (Li₂CO₃) were mixed together to a molar ratio of 2: 2: 1 for thirty minutes in a mortar to give a precursor. The resulting precursor was calcined at 300 °C for 12 hours in a nitrogen atmosphere and further sintered in the nitrogen atmosphere at 600 °C for 24

hours to synthesize LiFePO₄. Using the so-prepared LiFePO₄ as the positive electrode active material, a test cell was prepared as in Example 1.

Of the LiFePO₄, synthesized as described above, measurement was made of the powder X-ray diffraction pattern under the above-mentioned measurement conditions.

Fig.8 shows the powder X-ray diffraction pattern of LiFePO₄, synthesized in Example 3. It is seen from Fig.8 that no impurity other than LiFePO₄ was confirmed to exist in the product so that the single-phase LiFePO₄ has been produced.

Also, from the powder X-ray diffraction pattern of LiFePO₄, synthesized in Examples 4 to 6, no impurity other than LiFePO₄ was confirmed to exist in the product, so that, in these Examples, the single-phase LiFePO₄ has been produced.

The powder X-ray diffraction pattern of LiFePO₄, synthesized in Comparative Example 2, is shown in Fig.9, from which it is seen that impurities other than LiFePO₄ are present in the product such that no single-phase LiFePO₄ has been yielded. Meanwhile, in Fig.3, the peak attributable to the impurity is marked with an asterisk *. This impurity is presumably a trivalent iron compound, such as Fe₂O₃.

The test cells, prepared as described above, were subjected to the charging/discharging test, using the same method as stated in Experiment 1.

Of the battery of Example 3, the charging/discharging characteristics for the first to third cycles are shown in Fig.10, from which it is seen that the battery of Example 3 showed a flat potential in the vicinity of 3.4 V, thus indicating that the battery has a high reversible charging/discharging capacity of 163 mAh/g, which is

close to the theoretical capacity of 170 mAh/g.

The relation between the number of cycles and the charging/discharging capacity of the battery of Example 3 is shown in Fig.11, from which it is seen that the battery of the Example 3 undergoes cyclic deterioration as low as 0.1%/cycle, thus demonstrating stable battery characteristics.

Similarly to the battery of Example 3, the batteries of Examples 4 to 6 were confirmed to exhibit a high charging/discharging capacity and superior cyclic characteristics.

Test cells were prepared in the same way as in Example 6 except changing the sintering temperature of the precursor to check into the charging/discharging capacity of the cells. The relation between the sintering temperature of the precursor and the charging/discharging capacity of the cells is shown in Fig.12, from which it is seen that, by synthesizing LiFePO₄ at a temperature not lower than 350°C and not higher than 790°C, a high capacity exceeding 120 mAh/g of a conventional non-aqueous electrolyte secondary battery can be achieved.

The charging/discharging characteristics of the battery of Comparative Example 2 are shown in Fig.13. The relation between the number of cycles and the charging/discharging capacity of the battery of Comparative Example 2 is shown in Fig.14. It is seen from Figs.13 and 14 that the battery of Comparative Example 2 is satisfactory in the cyclic characteristics, however, it is high in capacity loss in an initial state and low in the charging/discharging efficiency.

Thus, it is seen that, by de-aerating the precursor in synthesizing LiFePO₄, it is possible to prevent Fe²⁺ in the precursor from being oxidized with residual oxygen to Fe³⁺ to yield single-phase LiFePO₄. It is also seen that the battery employing this single-phase LiFePO₄ as the positive electrode active material is superior in charging/discharging characteristics and cyclic characteristics.

<Experiment 3-1>

In an experiment 3-1, a variety of positive electrode active materials, containing LiFePO₄ or LiFe_{0.5}Mn_{0.5}PO₄, were synthesized, and non-aqueous electrolyte secondary batteries, employing these positive electrode active materials, were prepared, and evaluation was made of the difference in characteristics of these non-aqueous electrolyte secondary batteries.

Example 7

An LiFePO₄ carbon compound material composed of LiFePO₄ as a positive electrode active material and carbon as an electrically conductive material was synthesized and, using this compound material, non-aqueous electrolyte secondary batteries were prepared.

First, iron oxalate dihydride (FeC₂O₄·2H₂O), ammonium dihydrogen phosphate (NH₄H₂PO₄) and lithium carbonate (Li₂CO₃) were mixed to a molar ratio of 2: 2: 1 and thoroughly pulverized and mixed by a ball mill. The resulting mixture was calcined in a nitrogen atmosphere at 300°C for two hours to yield a precursor. This precursor and graphite were mixed to a weight ratio of 90:10 and pulverized sufficiently by a ball

mill for mixing. The precursor then was sintered in a nitrogen atmosphere at 550 °C for 24 hours to synthesize an LiFePO₄ carbon compound material.

A battery was prepared, using the LiFePO₄ carbon compound material, obtained as described above, as a positive electrode active material first, 97 wt% of the dried LiFePO₄ carbon compound material and 3 wt% of polyvinylidene fluoride, as a binder, were evenly mixed into N-methyl pyrrolidone as a solvent to prepare a paste-like positive electrode mixture. Meanwhile, #1300 manufactured by Aldrich Inc. was used as the polyvinylidene fluoride. This positive electrode mixture was applied to an aluminum mesh, as a current collector, and dried in situ in a dry argon atmosphere at 100 °C for one hour to form a layer of the positive electrode active material. The aluminum mesh, on which the layer of the positive electrode active material was formed, was punched to a disc 15.5 mm in diameter to form a pellet-like positive electrode. Meanwhile, this positive electrode carries 60 mg of the active material. A metal lithium foil was punched to substantially the same shape as the positive electrode and used as a negative electrode. In a mixed solvent of equal parts in volume of propylene carbonate and dimethyl carbonate was dissolved LiPF₆ at a concentration of 1 mol/l to prepare a non-aqueous electrolytic solution.

The positive electrode, prepared as described above, was accommodated in the positive electrode can, whilst the negative electrode was accommodated in the negative electrode can and a separator was arranged between the positive electrode and the negative electrode. The non-aqueous electrolytic solution was charged into the

positive electrode can and the negative electrode can. The electrode cans 3, 5 are caulked fixedly through the insulating gasket 7 to complete a 2025 type coin-shaped non-aqueous electrolyte secondary battery as a test cell.

Example 8

An LiFePO₄ precursor was prepared in the same way as in Example 7. A LiFePO₄ carbon compound material was prepared in the same way as in Example 7 except mixing this precursor, graphite and acetylene black together to a weight ratio of 90: 5: 5 and sufficiently pulverizing the mixture by a ball mill for mixing.

Using the LiFePO₄ carbon compound material, obtained as described above, as a positive electrode active material, a test cell was prepared in the same way as in Example 7.

Example 9

A LiFe_{0.5}Mn_{0.5}PO₄ carbon compound material was prepared as the positive electrode active material and a test cell was prepared using the LiFe_{0.5}Mn_{0.5}PO₄ carbon compound material as the positive electrode active material.

First, iron oxalate dihydride (FeC₂O₄·2H₂O), manganese acetate tetrahydride (Mn(CH₃OO)₂), ammonium dihydrogen phosphate (NH₄H₂PO₄) and lithium carbonate (Li₂CO₃) were mixed together to a molar ratio of 1: 1: 2: 1. The resulting mixture was sufficiently pulverized by a ball mill for mixing. The resulting mixture was calcined in a nitrogen atmosphere at 300 °C for two hours give a precursor. This precursor was mixed with graphite and acetylene black to give a weight ratio of 90 : 6: 4 and

sufficiently pulverized by a ball mill for mixing. The so-calcined precursor was sintered in the nitrogen atmosphere at 450°C for 24 hours to synthesize the $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$ carbon compound material.

Using the $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$ carbon compound material, obtained as described above, as a positive electrode active material, a test cell was prepared in the same way as in Example 7.

Comparative Example 3

As Comparative Example 3, LiFePO_4 was synthesized as a positive electrode active material and a test cell was fabricated using LiFePO_4 as the positive electrode active material.

First, iron oxalate dihydride ($\text{FeC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$), ammonium dihydrogen phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$) and lithium carbonate (Li_2CO_3) were mixed to a molar ratio of 2: 2: 1 and thoroughly pulverized and mixed by a ball mill. The resulting mixture was calcined in a nitrogen atmosphere at 300°C for two hours to yield a precursor. This precursor was sintered in a nitrogen atmosphere at 550°C for 24 hours to synthesize LiFePO_4 .

A battery was prepared, using the LiFePO_4 , obtained as described above, as a positive electrode active material. First, 87 wt% of the dried LiFePO_4 , 10 wt% of graphite, as an electrically conductive agent, and 3 wt% of polyvinylidene fluoride, as a binder, were evenly mixed into N-methyl pyrrolidone as a solvent to prepare a paste-like positive electrode mixture. Meanwhile, #1300 manufactured by Aldrich Inc. was used as the polyvinylidene fluoride. This positive electrode mixture was applied to an

aluminum mesh, as a current collector, and dried in situ in a dry argon atmosphere at 100°C for one hour to form a layer of the positive electrode active material. The aluminum mesh, on which the layer of the positive electrode active material was formed, was punched to a disc 15.5 mm in diameter to form a pellet-like positive electrode. Meanwhile, this positive electrode carries 60 mg of the active material. A metal lithium foil was punched to substantially the same shape as the positive electrode and used as a negative electrode. In a mixed solvent of equal parts in volume of propylene carbonate and dimethyl carbonate was dissolved LiPF₆ at a concentration of 1 mol/l to prepare a non-aqueous electrolytic solution.

The positive electrode, prepared as described above, was accommodated in the positive electrode can, whilst the negative electrode was accommodated in the negative electrode can and a separator was arranged between the positive electrode and the negative electrode. The non-aqueous electrolytic solution was charged into the positive electrode can and the negative electrode can. The positive and negative electrode cans were caulked fixedly through the insulating gasket 7 to complete a 2025 type coin-shaped non-aqueous electrolyte secondary battery.

Of the compound samples, synthesized in the Examples 7 to 9, and LiFePO₄, prepared in the Comparative Example 3, the powder X-ray diffraction pattern was measured under the same measurement conditions shown in Experiment 1 above. The measured results of the powder X-ray diffraction are shown in Figs.15 to 18.

The powder X-ray diffraction pattern of LiFePO₄, synthesized in Example 7, is

shown in Fig.15, from which it is seen that no impurities other than LiFePO_4 are confirmed in the product such that the product yielded is single-phase LiFePO_4 . Meanwhile, in Fig.3, the peak attributable to the impurity is marked with an asterisk *.

The precursor was sintered at plural temperature points in a range from 320°C to 850°C to prepare the LiFePO_4 carbon compound material in the same manner as in Example 7. It was found that there was no impurity present in the LiFePO_4 carbon compound material thus testifying to the yielding of the single-phase LiFePO_4 .

The powder X-ray diffraction pattern of the LiFePO_4 carbon compound material synthesized in Example 8 is shown in Fig.16, from which it is seen that the presence of impurities other than LiFePO_4 is not confirmed in the product, except the diffraction peak of graphite at approximately 26°, thus testifying to the yielding of the single-phase LiFePO_4 . In Fig.16, the peak attributable to graphite is indicated by an asterisk *.

The powder X-ray diffraction pattern of the $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$ carbon compound material, synthesized in Example 9, is shown in Fig.9. The powder X-ray diffraction patterns of the $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$ carbon compound materials, synthesized with the sintering temperatures of 500°C, 600°C and 700°C, are also shown in Fig.17. It may be seen from Fig.17 that, in the $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$ carbon compound materials, synthesized at several temperature points in the range from 450°C to 850°C, the presence of impurities other than LiFePO_4 is not confirmed, except the diffraction peak

of graphite at approximately 26°, thus testifying to the yielding of the single-phase LiFePO₄. In Fig.16, the peak attributable to graphite is indicated by an asterisk *.

The powder X-ray diffraction pattern of LiFePO₄, synthesized in Comparative Example 3, is shown in Fig.18, from which it is seen that no impurities other than LiFePO₄ are confirmed to exist in the product such that the product yielded is single-phase LiFePO₄.

The following charging/discharging test was conducted on the test cells of the Examples 7 to 9 and the Comparative Example 3, prepared as described above, to evaluate battery characteristics.

As this charging/discharging test, each test cell was charged by constant current charging and, when the battery voltage reached 4.5V, the charging system was switched from the constant current charging to constant voltage charging, and charging was carried out as the voltage of 4.5 V was kept. The charging was stopped when the current fell below 0.05 mA/cm². The discharging then was carried out and stopped at a time point when the battery voltage was lowered to 2.0 V.

Both charging and discharging were carried out at ambient temperature (25°C) and, in evaluating charging/discharging characteristics, the current density was set to 0.12 mA/cm². In evaluating load characteristics, the charging/discharging test was conducted under various different current densities. The results of the above charging/discharging test are shown in Figs.19 to 23.

Of the battery of Example 7, the charging/discharging characteristics for the

first to third cycles are shown in Fig.19, from which it is seen that the battery of Example 7 showed a flat potential in the vicinity of 3.4 V, thus indicating that the battery has a high reversible charging/discharging capacity of 155 mAh/g, which is close to the theoretical capacity of 170 mAh/g.

The relation between the number of cycles and the charging/discharging capacity of the battery of Example 7 is shown in Fig.20, from which it is seen that the battery of the Example 7 undergoes cyclic deterioration as low as 0.1%/cycle, thus demonstrating stable battery characteristics.

The load characteristics of the battery of Example 8 are shown in Fig.21, from which it is seen that, if charged/discharged at 0.2 mA, the battery of Example 8 gives a high capacity of 163 mAh/g, and that the reversible charging/discharging capacity for the charging/discharging current density of 2 mA/cm² is kept at 93% of that for the charging/discharging current density of 0.2 mA/cm².

The charging/discharging characteristics of the battery of Example 9 are shown in Fig.22, from which it is seen that the battery of Example 9 has the reversible charging/discharging capacity of 146 mAh/g close to the theoretical capacity of 170 mAh/g, with the average discharging potential being 3.58 V.

The load characteristics of the batteries of Example 7 and Comparative Example 3 are shown in Fig.23, from which it is seen that the reversible charging/discharging capacity for the charging/discharging current density of 1 mA/cm² is kept at 90% of that for the charging/discharging current density of 0.2

mA/cm². In the battery of the Comparative Example 3, the same amount of graphite as that used as the electrically conductive agent in Example 7 is added as the electrically conductive agent in the positive electrode mixture. However, the battery of the Comparative Example 3 is low in battery capacity and inferior in load characteristics as compared to the battery of Example 7.

The charging/discharging curve of the battery employing $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$ solely as the positive electrode mixture according to the teaching of the publication J. Electroche. Soc. 144,1188 (1997), is shown in Fig.11. Meanwhile, the battery employing $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$ solely as the positive electrode mixture is the battery of Comparative Example 4. It is seen from Fig.24 that the battery of the Comparative Example 4 has the battery capacity was low as 75 mAh/g.

The above results indicate that the batteries of Examples 7 and 8 are non-aqueous electrolyte secondary batteries containing LiFPO_4 and carbon as the positive electrode active material and as the electrically conductive material, respectively, and hence are superior in load characteristics and of high capacity as compared to the battery of the Comparative Example 3 employing solely the compound of the general formula LiFePO_4 as the positive electrode mixture. On the other hand, the battery of Example 9 is the non-aqueous electrolyte secondary battery containing LiFePO_4 and carbon as the positive electrode active material and as the electrically conductive material, respectively, and hence is superior in load characteristics and of high capacity as compared to the battery of the Comparative Example 4 containing solely the

compound represented by the general formula LiFePO_4 and carbon as the compound sample as the positive electrode active material and as the electrically conductive agent, respectively.

It is also seen from Figs.23 and 21 that, on comparing Examples 7 and 8, the battery of Example 8 is superior in load characteristics to the battery of Fig.7 in which graphite alone is added to the precursor. From this it is seen that, with the non-aqueous electrolyte secondary battery, load characteristics may be improved by employing graphite and acetylene black in combination as carbon in synthesizing the positive electrode active material.

Comparison of the Examples 7 and 8 in Figs.19 and 22 also reveals that the battery of Example 9 has a higher value of the discharging potential. From this it is seen that, by employing the $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$ carbon compound material as the positive electrode active material, it is possible to realize a non-aqueous electrolyte secondary battery of higher capacity and higher voltage value.

<Experiment 3-2>

In the experiment 3-2, a variety of positive electrode active materials containing LiMnPO_4 were synthesized and non-aqueous electrolyte secondary batteries containing these positive electrode active materials were prepared, and evaluation was made of difference in characteristics of the respective non-aqueous electrolyte secondary batteries.

Example 10

As a positive electrode active material, LiFePO₄ carbon compound material was synthesized and a test cell containing the LiFePO₄ carbon compound material as the positive electrode active material was prepared.

First, manganese carbonate (MnCO₃), ammonium dihydrogen phosphate (NH₄H₂PO₄) and lithium carbonate (Li₂CO₃) were mixed to a molar ratio of 2: 2: 1 and thoroughly pulverized and mixed by a ball mill. The resulting mixture was calcined in a nitrogen atmosphere at 280°C for three hours to yield a precursor. This precursor and acetylene black were mixed to a weight ratio of 95:5 and pulverized sufficiently by a ball mill for mixing. The precursor then was sintered in a nitrogen atmosphere at 600°C for 24 hours to synthesize an LiFePO₄ carbon compound material.

A battery was prepared, using the LiFePO₄, obtained as described above, as a positive electrode active material. First, 97 wt% of the dried LiFePO₄ carbon compound material, 10 wt% of graphite, and 3 wt% of polyvinylidene fluoride, as a binder, were evenly mixed into N-methyl pyrrolidone as a solvent to prepare a paste-like positive electrode mixture. Meanwhile, #1300 manufactured by Aldrich Inc. was used as the polyvinylidene fluoride. This positive electrode mixture was applied to an aluminum mesh, as a current collector, and dried in situ in a dry argon atmosphere at 100°C for one hour to form a layer of the positive electrode active material. The aluminum mesh, on which the layer of the positive electrode active material was formed, was punched to a disc 15.5 mm in diameter to form a pellet-like positive electrode. Meanwhile, this positive electrode carries 60 mg of the active material. A

metal lithium foil was punched to substantially the same shape as the positive electrode and used as a negative electrode. In a mixed solvent of equal parts in volume of propylene carbonate and dimethyl carbonate was dissolved LiPF₆ at a concentration of 1 mol/l to prepare a non-aqueous electrolytic solution.

The positive electrode, prepared as described above, was accommodated in the positive electrode can, whilst the negative electrode was accommodated in the negative electrode can and a separator was arranged between the positive electrode and the negative electrode. The non-aqueous electrolytic solution was charged into the positive electrode can and the negative electrode can. The positive and negative electrode cans were caulked fixedly through the insulating gasket 7 to complete a 2025 type coin-shaped non-aqueous electrolyte secondary battery.

Example 11

An LiFePO₄ carbon compound material was prepared in the same way as in Example 10 except mixing the precursor and acetylene black in a weight ratio of 90:10. Using this LiFePO₄ carbon compound material as the positive electrode mixture, a test cell was prepared in the same way as in Example 10.

Example 12

The LiFePO₄ carbon compound material was synthesized in the same way as in Example 10 to synthesize LiFePO₄ carbon compound material. Using this LiFePO₄ carbon compound material as the positive electrode mixture, a test cell was prepared in the same way as in Example 10.

Comparative Example 5

As a positive electrode active material, LiMnPO_4 was synthesized and a test cell containing the LiFePO_4 carbon compound material as the positive electrode active material was prepared.

First, manganese carbonate (MnCO_3), ammonium hydrogen phosphate ($\text{NH}_4\text{H}_2\text{PO}_4$) and lithium carbonate (Li_2CO_3) were mixed to a molar ratio of 2: 2: 1 and thoroughly pulverized and mixed by a ball mill. The resulting mixture was calcined in a nitrogen atmosphere at 300°C for three hours to yield a precursor. This precursor then was sintered in a nitrogen atmosphere at 600°C for 24 hours to synthesize LiMnPO_4 .

For preparing a battery employing LiMnPO_4 , obtained as described above, as the positive electrode active material, a test cell was prepared in the same way as in Example 10, except evenly mixing 85 wt% of dried LiMnPO_4 as a positive electrode mixture, 10 wt% of graphite as an electrically conductive agent and 10 wt% of polyvinylidene fluoride as a binder, in N-methyl pyrrolidone, as a solvent, to prepare a paste-like positive electrode mixture.

Of the compound samples of Examples 10 to 12, synthesized as described above, and LiMnPO_4 of Comparative Example 5, a powder X-ray diffraction pattern was measured in accordance with the method and conditions described above.

The powder X-ray diffraction pattern of the LiMnPO_4 carbon compound material, synthesized in accordance with Examples 10 to 12, is shown in Fig.25, from

which it is seen that no impurities other than LiMnPO₄ is confirmed to exist in the product, such that the product yielded is single-phase LiFePO₄.

The powder X-ray diffraction pattern of the LiMnPO₄ carbon compound material, synthesized in accordance with Comparative Example 5, is shown in Fig.26, from which it is seen that no impurities other than LiMnPO₄ is confirmed to exist in the product such that the product yielded is single-phase LiFePO₄.

The following charging/discharging test was conducted on the test cells of the Example 11 and the Comparative Example 5, prepared as described above, to evaluate battery characteristics. As this charging/discharging test, each test cell was charged by constant current charging with the current of 0.5 mA/cm². When the battery voltage reached 4.5V, the charging system was switched from the constant current charging to constant voltage charging, and charging was carried out as the voltage of 4.5 V was kept. The charging was stopped when the current fell below 0.05 mA/cm². The discharging then was carried out and stopped at a time point when the battery voltage was lowered to 2.0 V.

The following charging/discharging test was conducted on the test cell of the Example 12, prepared as described above, to evaluate battery characteristics. As this charging/discharging test, the test cell was charged by constant current charging with the current of 0.5 mA/cm². When the battery voltage reached 4.5V, the charging system was switched from the constant current charging to constant voltage charging, and charging was carried out as the voltage of 4.8 V was kept. The charging was

stopped when the current fell below 0.05 mA/cm^2 . The discharging then was carried out and stopped at a time point when the battery voltage was lowered to 2.0 V. Meanwhile, the battery charging and discharging were carried out at an ambient temperature (23°C). The results obtained by this charging/discharging test are shown in Figs.27 to 29.

Fig.27 shows a graph representing charging/discharging characteristics for the battery of Example 11. It is seen from Fig.27 that the battery of Fig.11 has a flat potential in the vicinity of 4V and generates the reversible charging/discharging capacity of 113 mAh/g . On the other hand, the charging/discharging characteristics of the battery of Example 12 are shown in a graph of Fig.28, from which it is seen that the discharging voltage is high, with the discharging capacity being as large as 120 mAh/g . Conversely, the battery of Comparative Example 5 is free of a flat discharging area, with there occurring no Mn oxidation.

It may be seen from the above results that the batteries of Examples 11 and 12 are non-aqueous electrolyte secondary batteries obtained by using a synthesized compound sample composed of LiMnPO_4 as a positive electrode mixture and of acetylene black as an electrically conductive agent, such that these batteries are superior in load characteristics and of high capacity as compared to the battery of Comparative Example 5 employing LiFePO_4 as the positive electrode active material.

For measuring the volume grain size distribution of the LiMnPO_4 carbon compound material of Examples 10 to 12 and LiMnPO_4 of Comparative Example 5,

the scattering of the laser light was measured using a volume grain size distribution measurement device, manufactured by HORIBA SEISAKUSHO CO. LTD. under the trade name of Micro-Lack grain size analyzer LA-920. The measured results of the volume grain size distribution are shown in Fig.30, from which is seen that the LiMnPO_4 carbon compound material and LiMnPO_4 all contain particles not larger than $10 \mu\text{m}$.

By representing the volume grain size distribution of the LiMnPO_4 carbon compound material of Example 12 by integrated amount of passed grains, it has been seen that the grain size in its entirety is not more than $6 \mu\text{m}$. That is, by adding carbon to the synthetic material of LiMnPO_4 or to the precursor, the grain growth of the precursor during the sintering process can be suppressed to realize more homogeneous finer compound sample.

Industrial Applicability

As may be seen from the foregoing description, the manufacturing method of the positive electrode mixture according to the present invention adds a reducing agent to the precursor during the mixing process. Since this prevents the 3d transition metal M in the synthetic starting material from being oxidized by residual oxygen, thus yielding a single-phase $\text{Li}_x\text{M}_y\text{PO}_4$ free of impurities, it is possible to prepare a positive electrode active material capable of reversibly and satisfactorily doping/undoping lithium.

Moreover, in the manufacturing method for the non-aqueous electrolyte secondary battery according to the present invention, a reducing agent is added to the precursor during the mixing process. This prevents the 3d transition metal M in the synthetic starting material from being oxidized by residual oxygen to yield a single-phase $\text{Li}_x\text{M}_y\text{PO}_4$ free of impurities. This $\text{Li}_x\text{M}_y\text{PO}_4$ renders it possible to produce a non-aqueous electrolyte secondary battery of high capacity because lithium is diffused sufficiently in the grains.

Moreover, with the manufacturing method of the positive electrode active material according to the present invention, air contained in the precursor is removed in the de-aerating process. This prevents the 3d transition metal M in the starting material for synthesis from being oxidized by residual oxygen to yield single-phase $\text{Li}_x\text{M}_y\text{PO}_4$ free of impurities, thus enabling the manufacture of a positive electrode active material capable of reversibly and satisfactorily doping/undoping lithium.

Moreover, with the manufacturing method of the positive electrode active material according to the present invention, air contained in the precursor is removed in the de-aerating process in preparing the positive electrode active material. This prevents the 3d transition metal M in the starting material for synthesis from being oxidized by residual oxygen to yield single-phase $\text{Li}_x\text{M}_y\text{PO}_4$ free of impurities. With this $\text{Li}_x\text{M}_y\text{PO}_4$, since lithium diffusion in the particle occurs sufficiently, it becomes possible to produce a high-capacity non-aqueous electrolyte secondary battery.

In addition, in the manufacturing method for the positive electrode active

material according to the present invention, an electrically conductive agent is added to the starting material for synthesis or to the precursor of the positive electrode active material. The positive electrode active material produced exhibits superior load characteristics and electrode molding performance thus realizing a high capacity.

Moreover, in the manufacturing method for the positive electrode active material according to the present invention, an electrically conductive agent is added to the starting material for synthesis or to the precursor in synthesizing the positive electrode active material. Since the positive electrode active material produced thus exhibits superior load characteristics and electrode molding performance, the non-aqueous electrolyte secondary battery produced exhibits high capacity as the non-aqueous electrolyte secondary battery.

Claims

1. A method for producing a positive electrode active material comprising:
 - a mixing step of mixing a plurality of substances to give a precursor, said substances proving a starting material for synthesis of a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M includes at least one of 3d transition metals; and
 - a sintering step of sintering and reacting said precursor obtained by said mixing step;
2. The method for producing a positive electrode active material according to claim 1 wherein said $\text{Li}_x\text{M}_y\text{PO}_4$ is LiFePO_4 .
3. The method for producing a positive electrode active material according to claim 2 wherein said reducing agent is Fe which is also a part of said starting material for synthesis.
4. A method for producing a non-aqueous electrolyte secondary battery having a positive electrode containing a positive electrode active material capable of reversibly doping/undoping lithium, a negative electrode mounted facing said positive electrode and capable of reversibly doping/undoping lithium, and a non-aqueous electrolyte interposed between said positive electrode and the negative electrode, said positive electrode active material being produced by a method comprising:
 - a mixing step of mixing a plurality of substances to give a precursor, said

substances proving a starting material for synthesis of a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M includes at least one of 3d transition metals; and

a sintering step of sintering and reacting said precursor obtained by said mixing step;

wherein a reducing agent is added to said precursor in said mixing step to prepare said positive electrode active material.

5. The method for producing a non-aqueous electrolyte secondary battery according to claim 4 wherein said $\text{Li}_x\text{M}_y\text{PO}_4$ is LiFePO_4 .
6. The method for producing a non-aqueous electrolyte secondary battery according to claim 5 wherein said reducing agent is Fe which is also a part of said starting material for synthesis.
7. A method for producing a positive electrode active material comprising:

a mixing step of mixing a plurality of substances to give a precursor, said substances proving a starting material for synthesis of a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M includes at least one of 3d transition metals;

a de-aerating step of removing air contained in said precursor obtained in said mixing step;

a sintering step of sintering and reacting said precursor obtained by said mixing step.

8. The method for producing a positive electrode active material according to claim 7 wherein an inert gas is introduced after substituting vacuum for an atmosphere in said precursor and subsequently an inert gas is introduced to remove air contained in said precursor.

9. The method for producing a positive electrode active material according to claim 7 wherein a solvent is allowed to co-exist with the precursor in said de-aerating step and is vaporized off in said inert gas to remove air contained in said precursor.

10. The method for producing a non-aqueous electrolyte secondary battery according to claim 7 wherein said $\text{Li}_x\text{M}_y\text{PO}_4$ is LiFePO_4 .

11. A method for producing a non-aqueous electrolyte secondary battery having a positive electrode containing a positive electrode active material capable of reversibly doping/undoping lithium, a negative electrode mounted facing said positive electrode and capable of reversibly doping/undoping lithium, and a non-aqueous electrolyte interposed between said positive electrode and the negative electrode, said positive electrode active material being produced by a method comprising:

a mixing step of mixing a plurality of substances to give a precursor, said substances proving a starting material for synthesis of a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M includes at least one of 3d transition metals;

a de-aerating step of de-aerating air contained in said precursor obtained in said mixing step; and

a sintering step of sintering and reacting said precursor obtained in a state free of air by said de-aerating step.

12. The method for producing a non-aqueous electrolyte secondary battery according to claim 11 wherein an inert gas is introduced after substituting vacuum for an atmosphere in said precursor and subsequently an inert gas is introduced to remove air contained in said precursor.

13. The method for producing a non-aqueous electrolyte secondary battery according to claim 11 wherein a solvent is allowed to co-exist with the precursor in said de-aerating step and is vaporized off in said inert gas to remove air contained in said precursor.

14. The method for producing a non-aqueous electrolyte secondary battery according to claim 11 wherein said $\text{Li}_x\text{M}_y\text{PO}_4$ is LiFePO_4 .

15. A method for producing a positive electrode active material comprising:

 a mixing step of mixing a plurality of substances to give a precursor, said substances proving a starting material for synthesis of a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M includes at least one of 3d transition metals; and

 a sintering step of sintering and reacting said precursor obtained by said mixing step;

 wherein an electrically conductive agent is added to said starting material for synthesis or said precursor.

16. The method for producing a positive electrode active material according to claim 15 wherein said electrically conductive agent is added in an amount of 0.5 to 20 parts by weight to 100 parts by weight of said $\text{Li}_x\text{M}_y\text{PO}_4$.

17. The method for producing a positive electrode active material according to claim 15 wherein said $\text{Li}_x\text{M}_y\text{PO}_4$ is LiFePO_4 .

18. The method for producing a positive electrode active material according to claim 15 wherein said $\text{Li}_x\text{M}_y\text{PO}_4$ is $\text{LiFe}_x\text{Mn}_{1-x}\text{PO}_4$.

19. The method for producing a positive electrode active material according to claim 15 wherein said $\text{Li}_x\text{M}_y\text{PO}_4$ is $\text{Li}_x\text{Mn}_y\text{PO}_4$ where x is such that $0 < x \leq 2$ and y is such that $0.8 \leq y \leq 1.2$.

20. The method for producing a positive electrode active material according to claim 15 wherein said electrically conductive agent is carbon.

21. A method for producing a non-aqueous electrolyte secondary battery having a positive electrode active material capable of reversibly doping/undoping lithium, a negative electrode mounted facing said positive electrode and capable of reversibly doping/undoping lithium, and a non-aqueous electrolyte interposed between said positive electrode and the negative electrode,

wherein said positive electrode active material is synthesized by a mixing step of mixing a plurality of substances to give a precursor, said substances proving a starting material for synthesis of a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M includes

at least one of 3d transition metals and a sintering step of sintering and reacting said precursor obtained by said mixing step, and wherein an electrically conductive agent is added to said starting material for synthesis or to said precursor to synthesize said positive electrode active material.

22. The method for producing a non-aqueous electrolyte secondary battery according to claim 21 wherein said electrically conductive agent is added in an amount of 0.5 to 20 parts by weight to 100 parts by weight of said $\text{Li}_x\text{M}_y\text{PO}_4$.

23. The method for producing a non-aqueous electrolyte secondary battery according to claim 21 wherein said $\text{Li}_x\text{M}_y\text{PO}_4$ is LiFePO_4 .

24. The method for producing a non-aqueous electrolyte secondary battery according to claim 21 wherein said $\text{Li}_x\text{M}_y\text{PO}_4$ is $\text{LiFe}_x\text{Mn}_{1-x}\text{PO}_4$.

25. The method for producing a non-aqueous electrolyte secondary battery according to claim 21 wherein said $\text{Li}_x\text{M}_y\text{PO}_4$ is $\text{Li}_x\text{Mn}_y\text{PO}_4$ where x is such that $0 < x \leq 2$ and y is such that $0.8 \leq y \leq 1.2$.

26. The method for producing a non-aqueous electrolyte secondary battery according to claim 21 wherein said electrically conductive agent is carbon.

ABSTRACT

In a method for producing a positive electrode active material including a mixing step of mixing a plurality of substances to give a precursor and a sintering step of sintering and reacting said precursor obtained by said mixing step, in which these substances prove a starting material for synthesis of a compound represented by the general formula $\text{Li}_x\text{M}_y\text{PO}_4$ where x is such that $0 < x \leq 2$, y is such that $0.8 \leq y \leq 1.2$ and M includes at least one of 3d transition metals,, a reducing agent is added in the above mixing step to said precursor to render it possible to prepare a positive electrode active material capable of reversibly and satisfactorily doping/undoping lithium.

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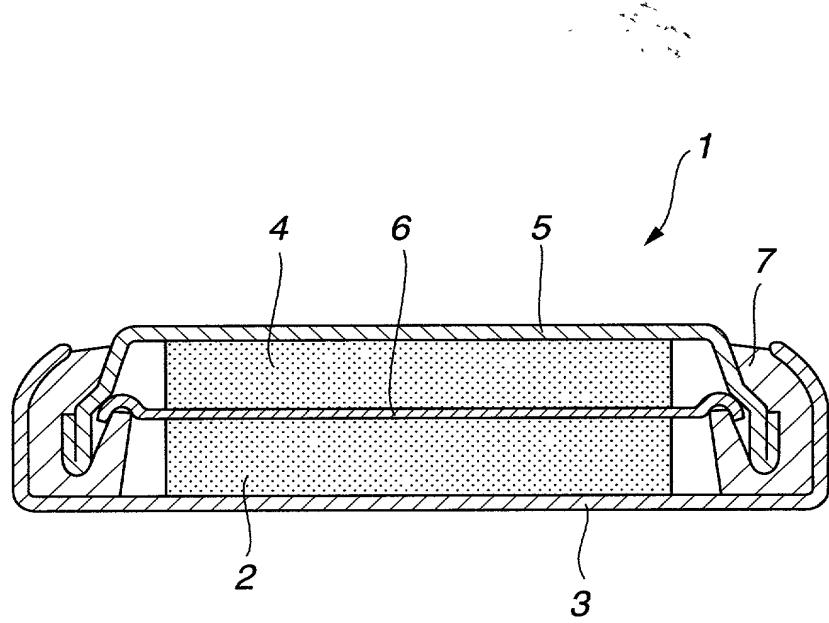


FIG.1

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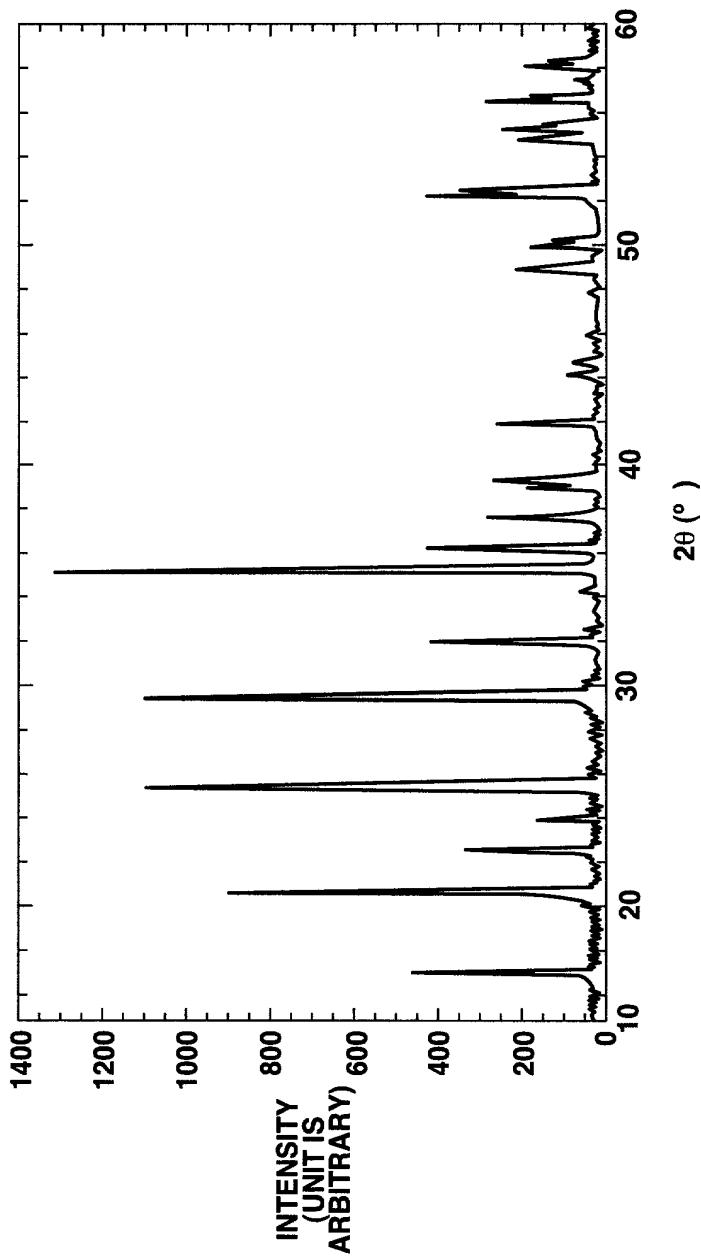


FIG.2

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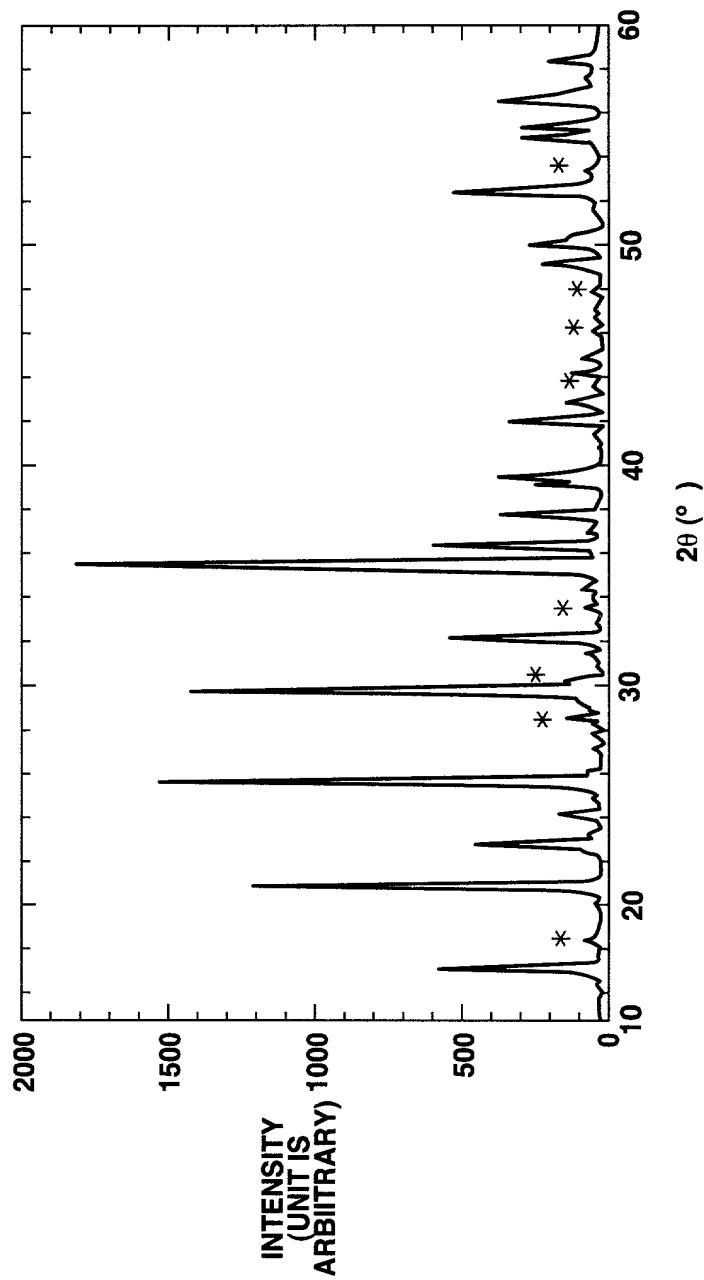
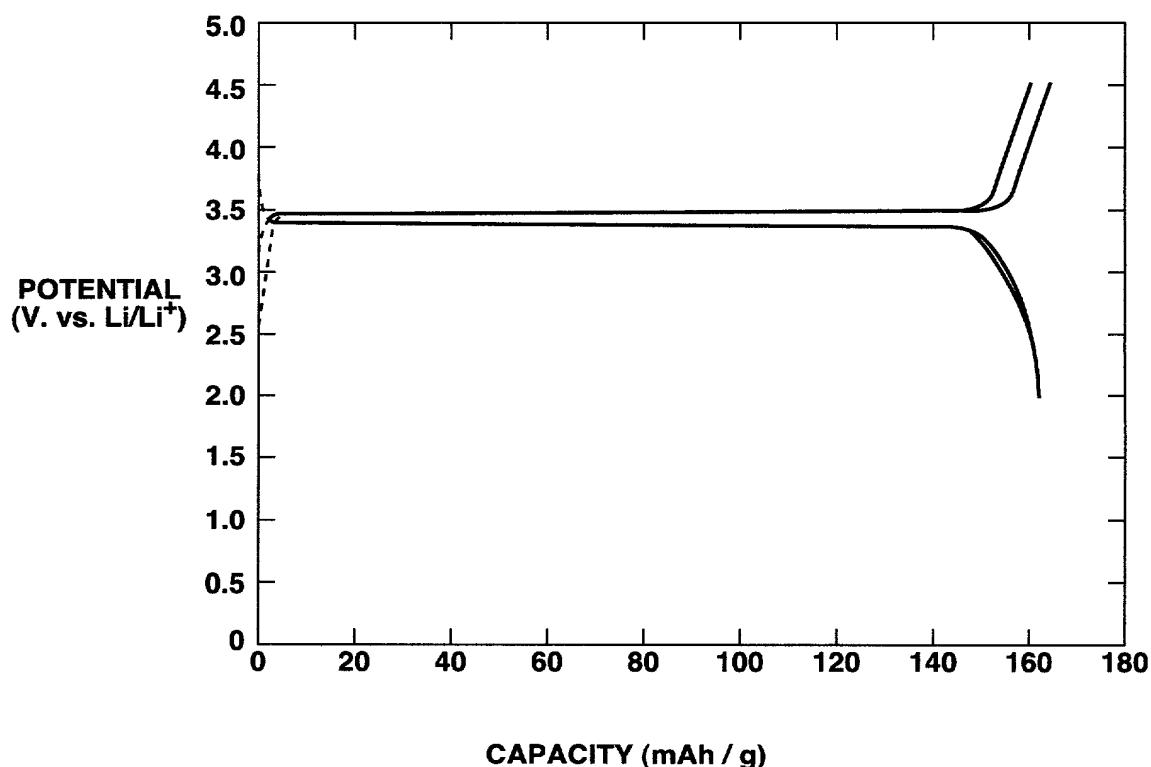


FIG.3

**FIG.4**

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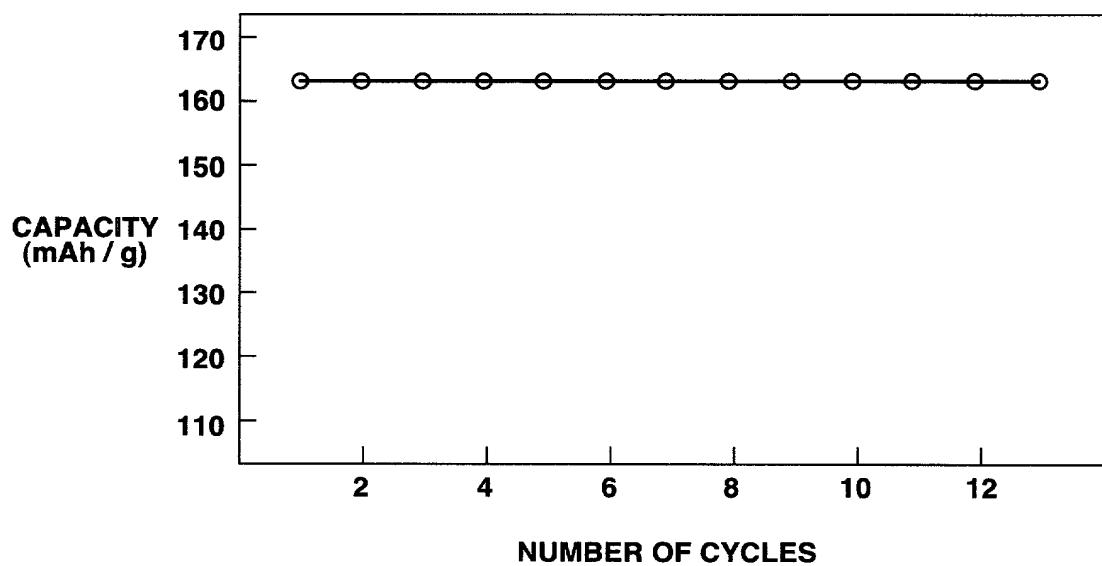


FIG.5

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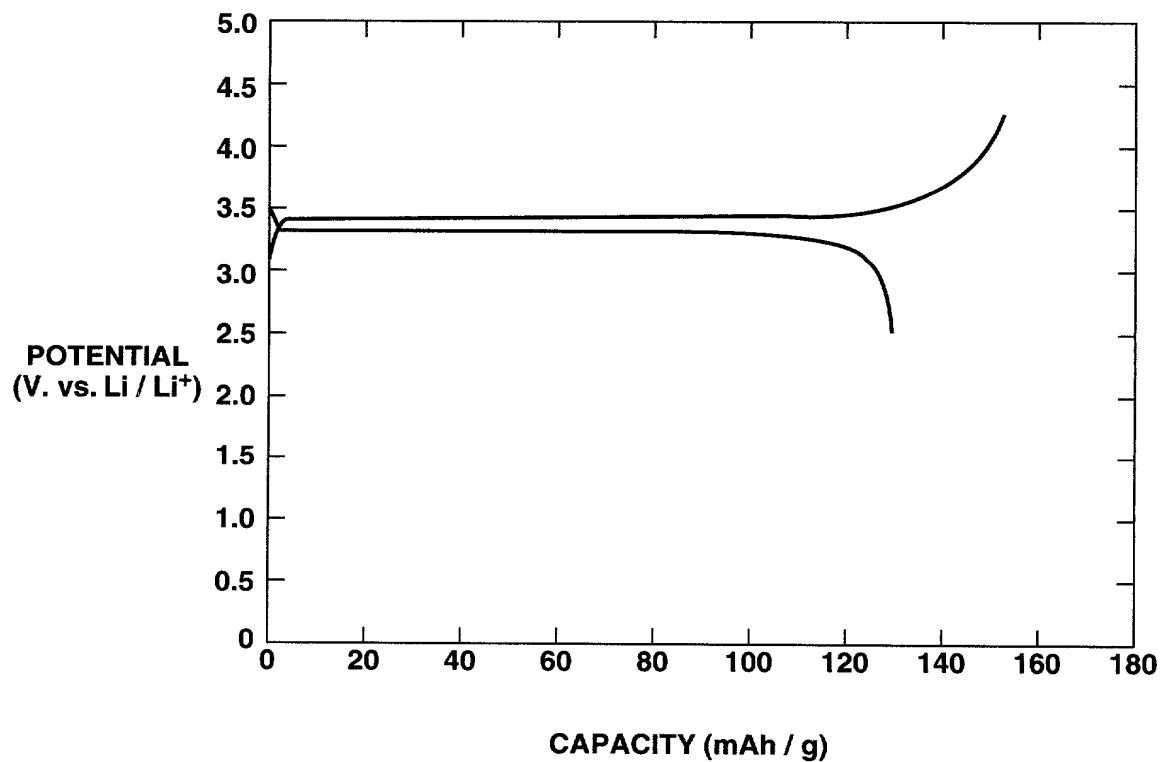


FIG.6

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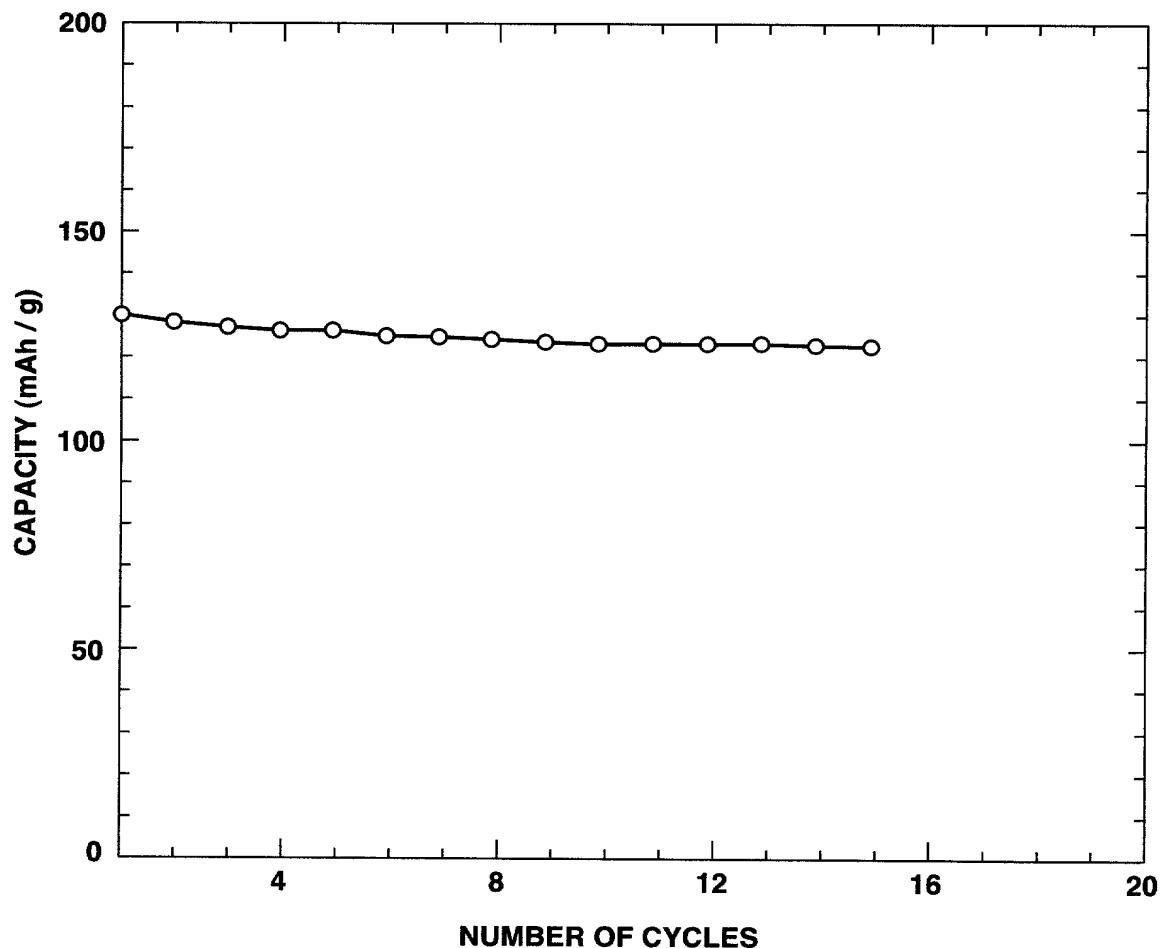


FIG.7

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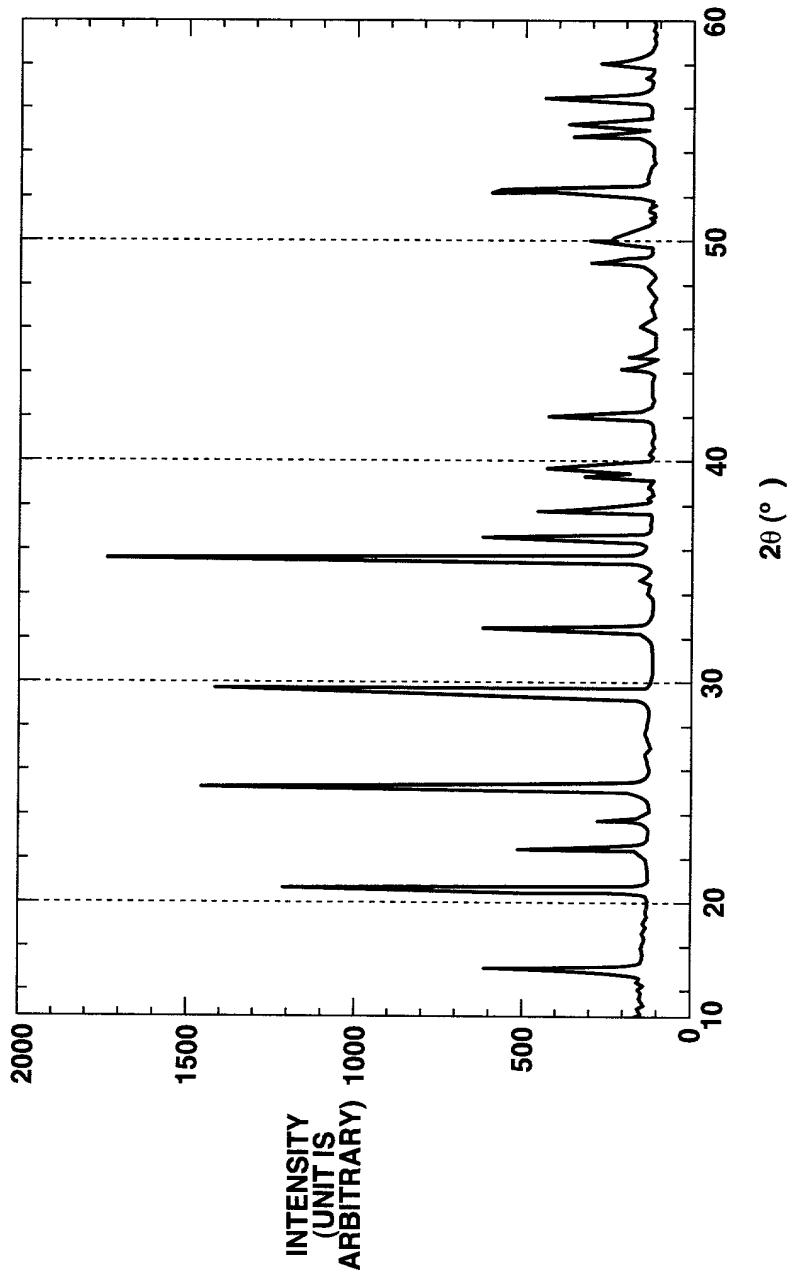


FIG.8

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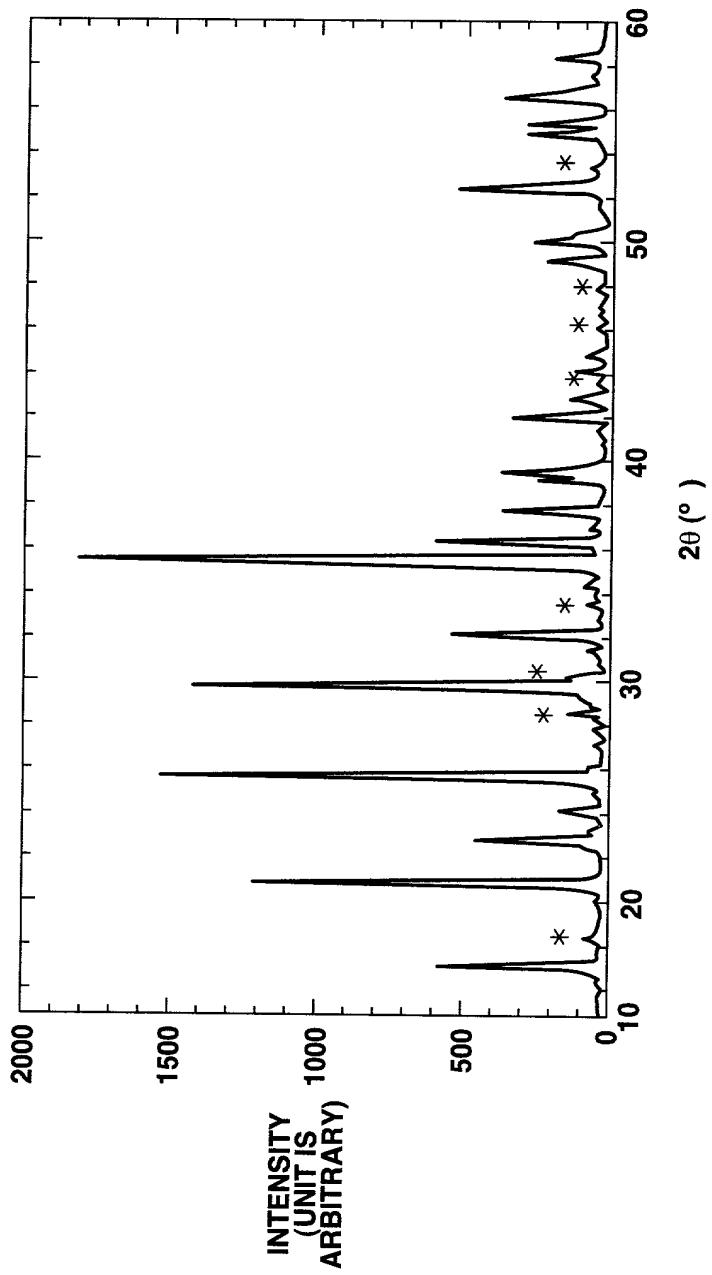


FIG.9

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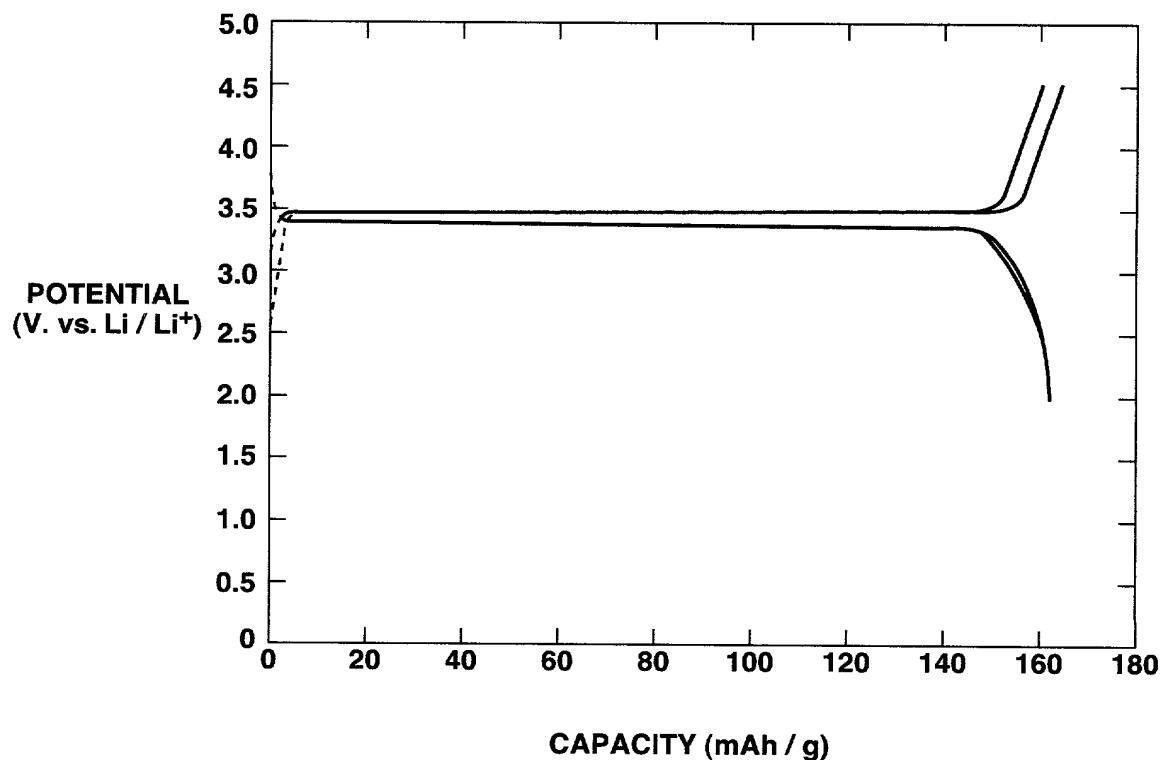


FIG.10

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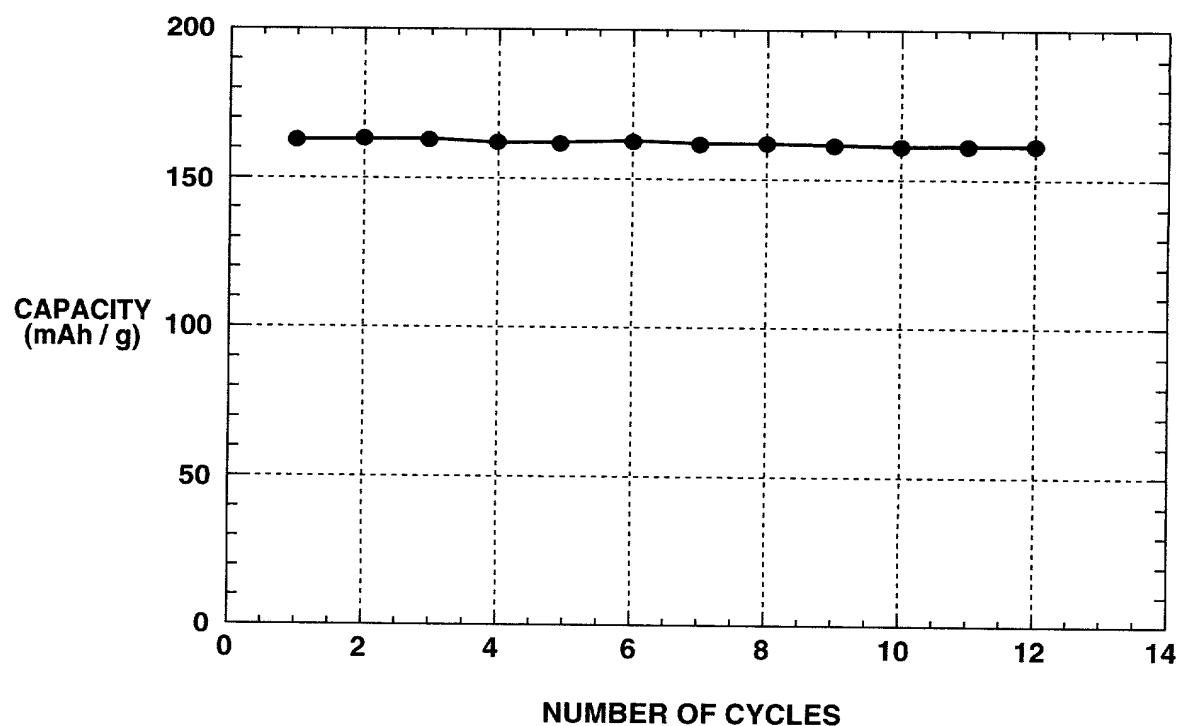


FIG.11

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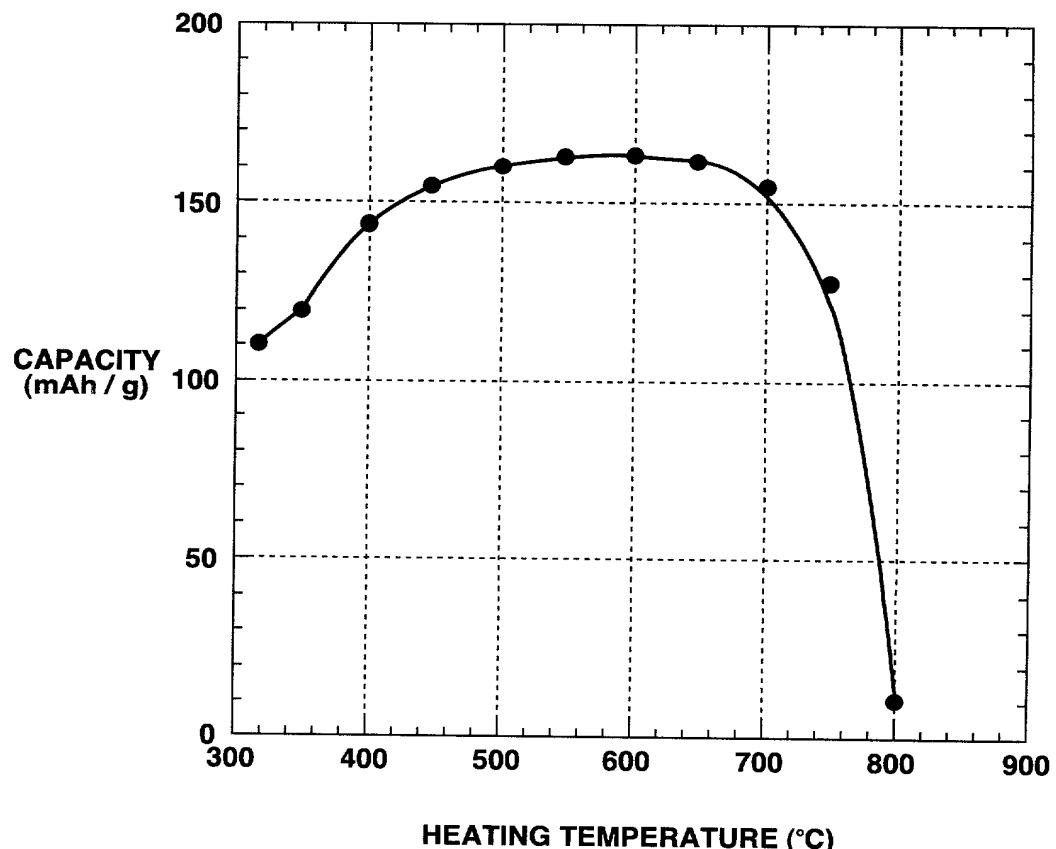


FIG.12

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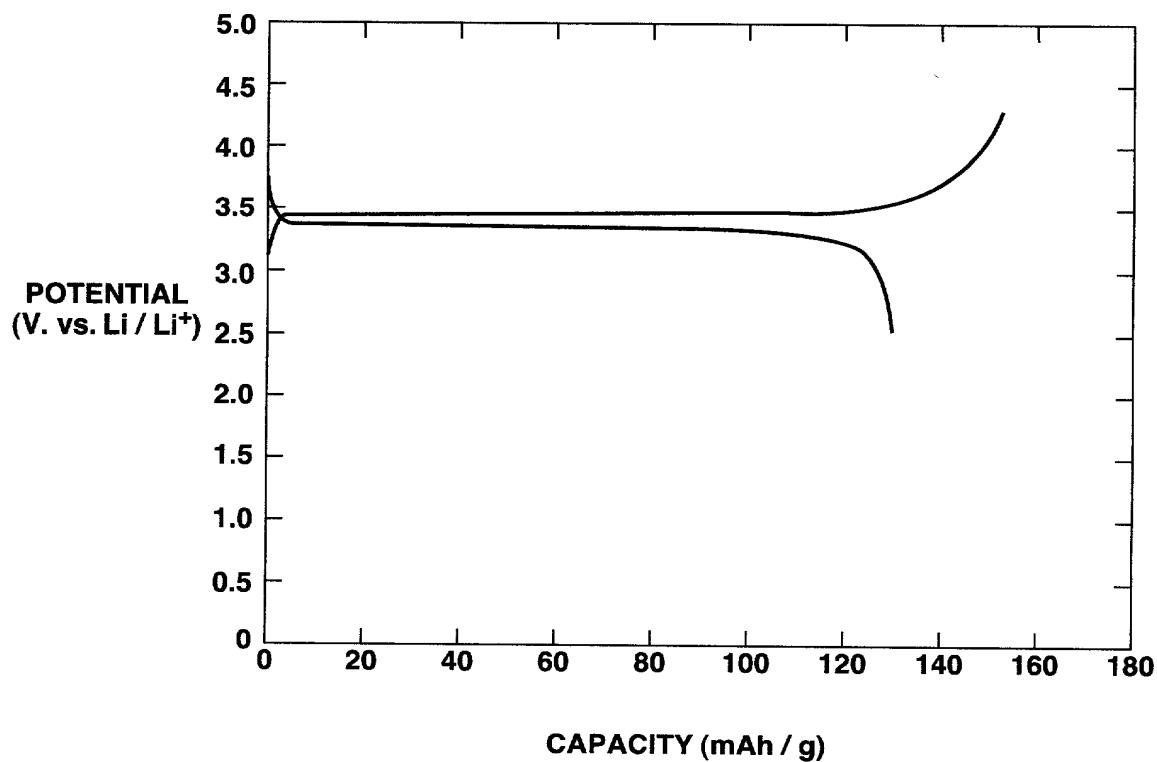


FIG.13

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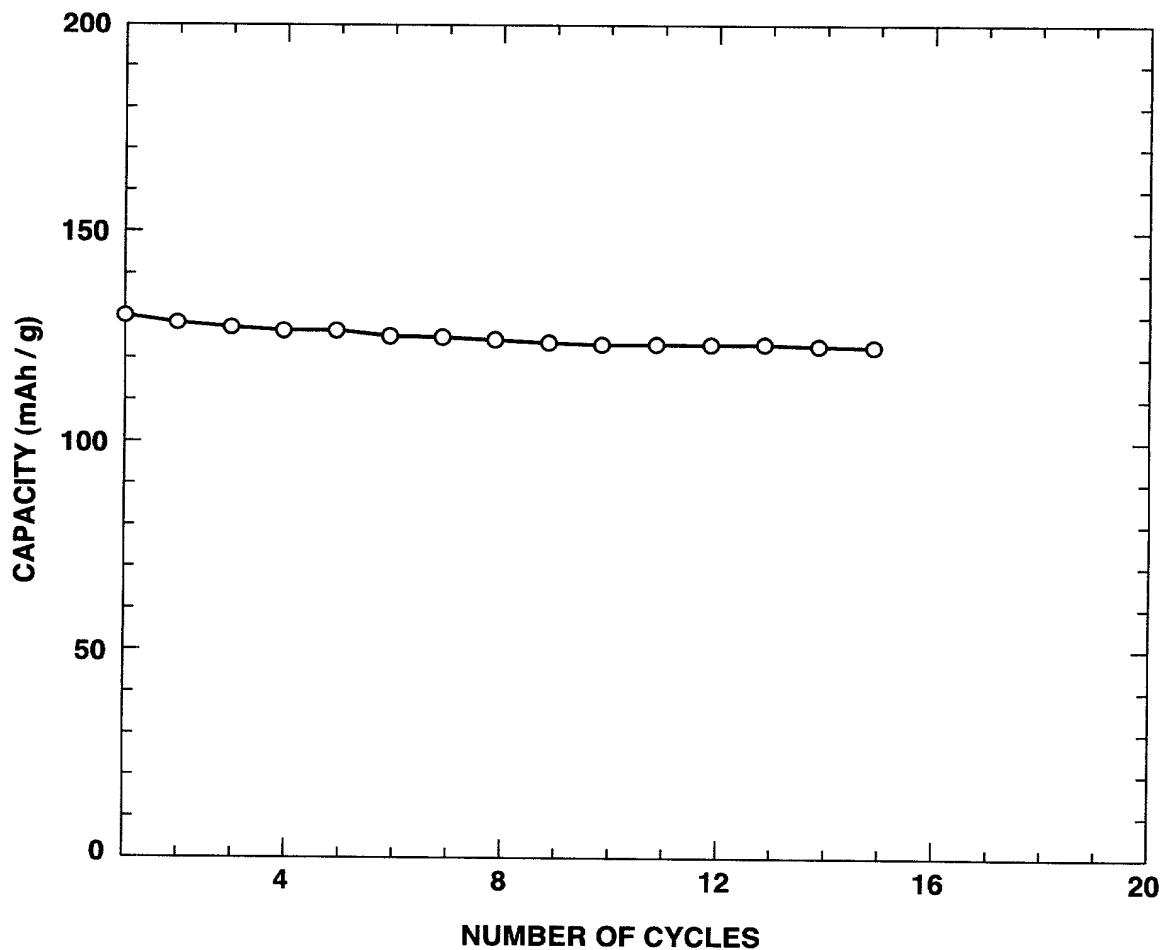


FIG.14

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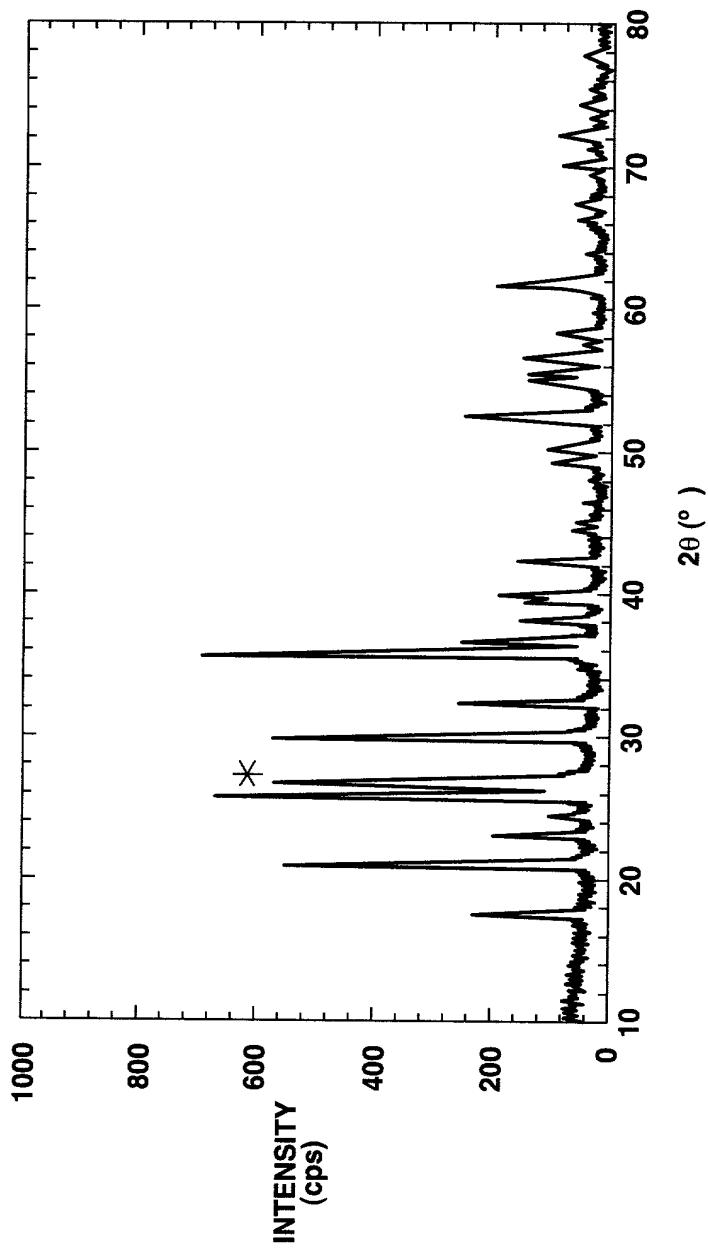


FIG.15

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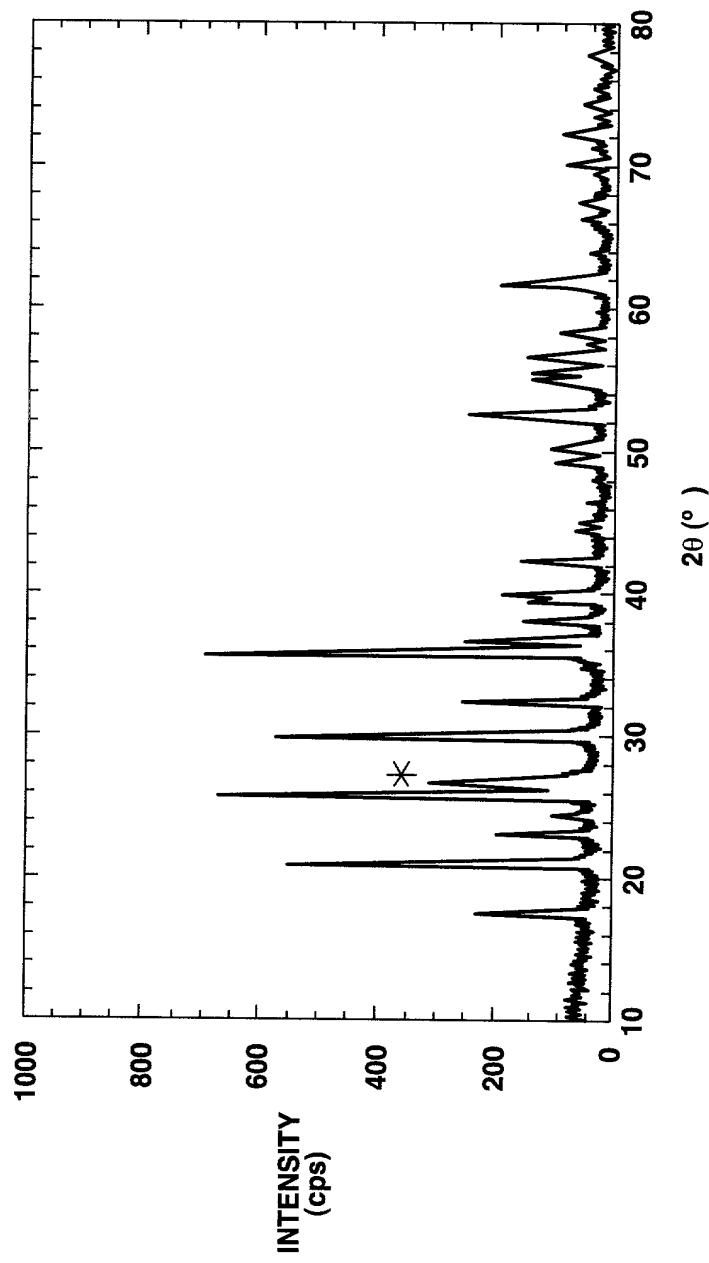


FIG. 16

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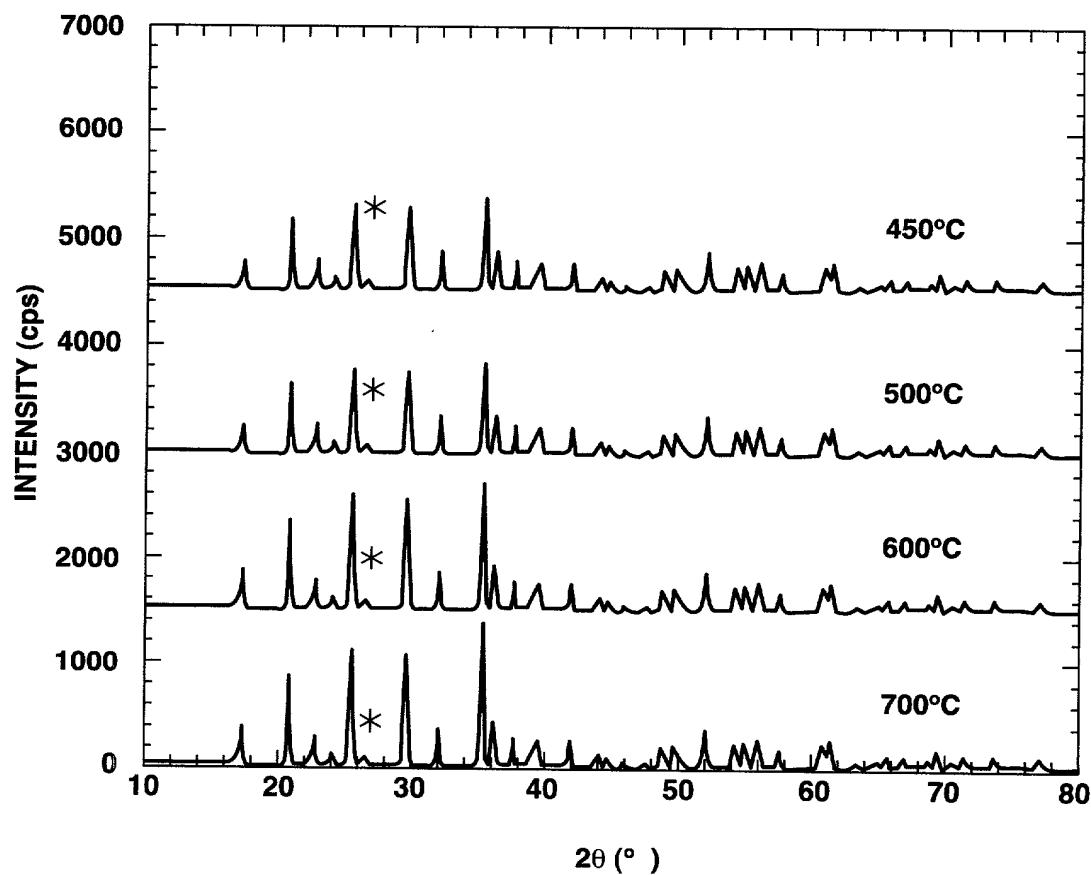


FIG.17

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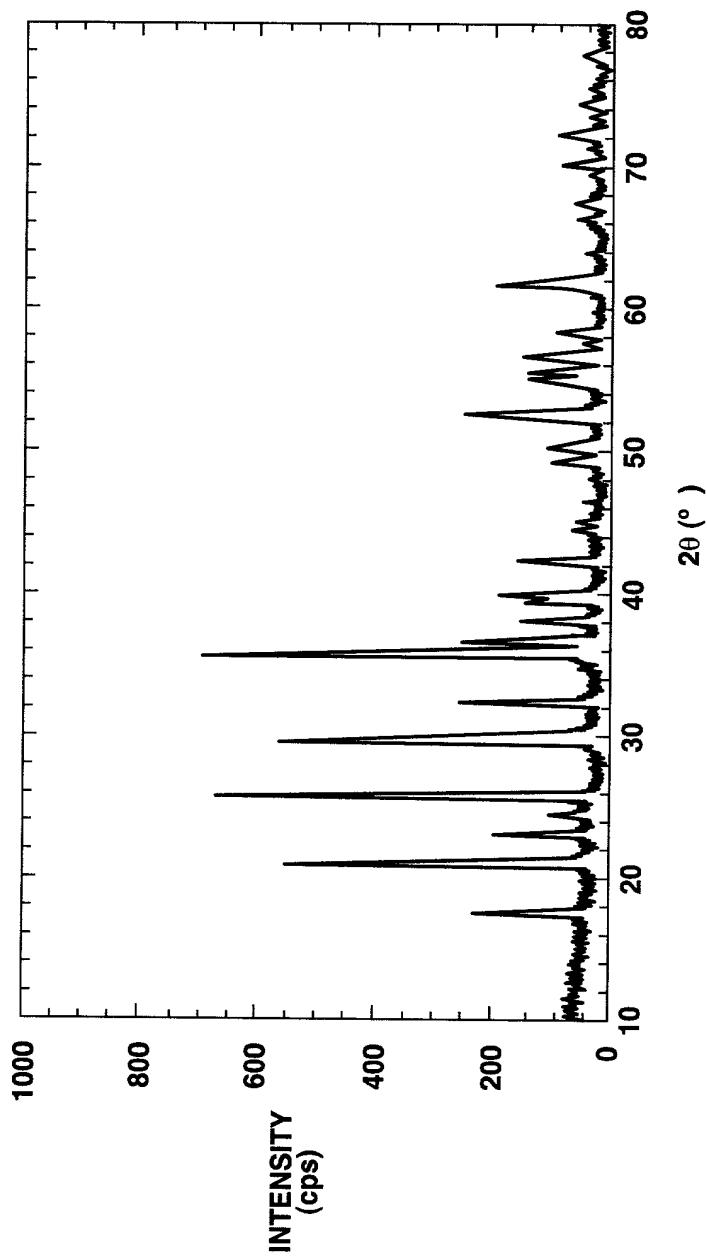


FIG.18

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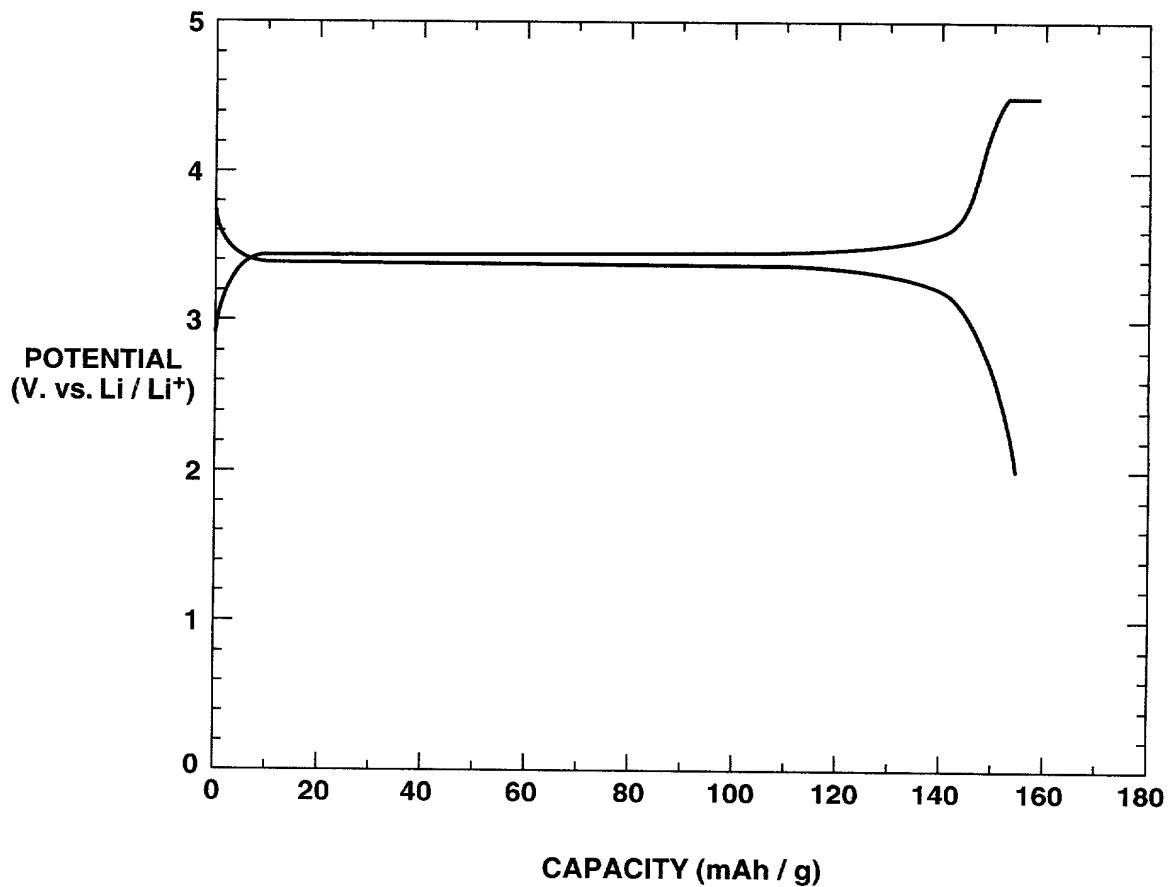


FIG.19

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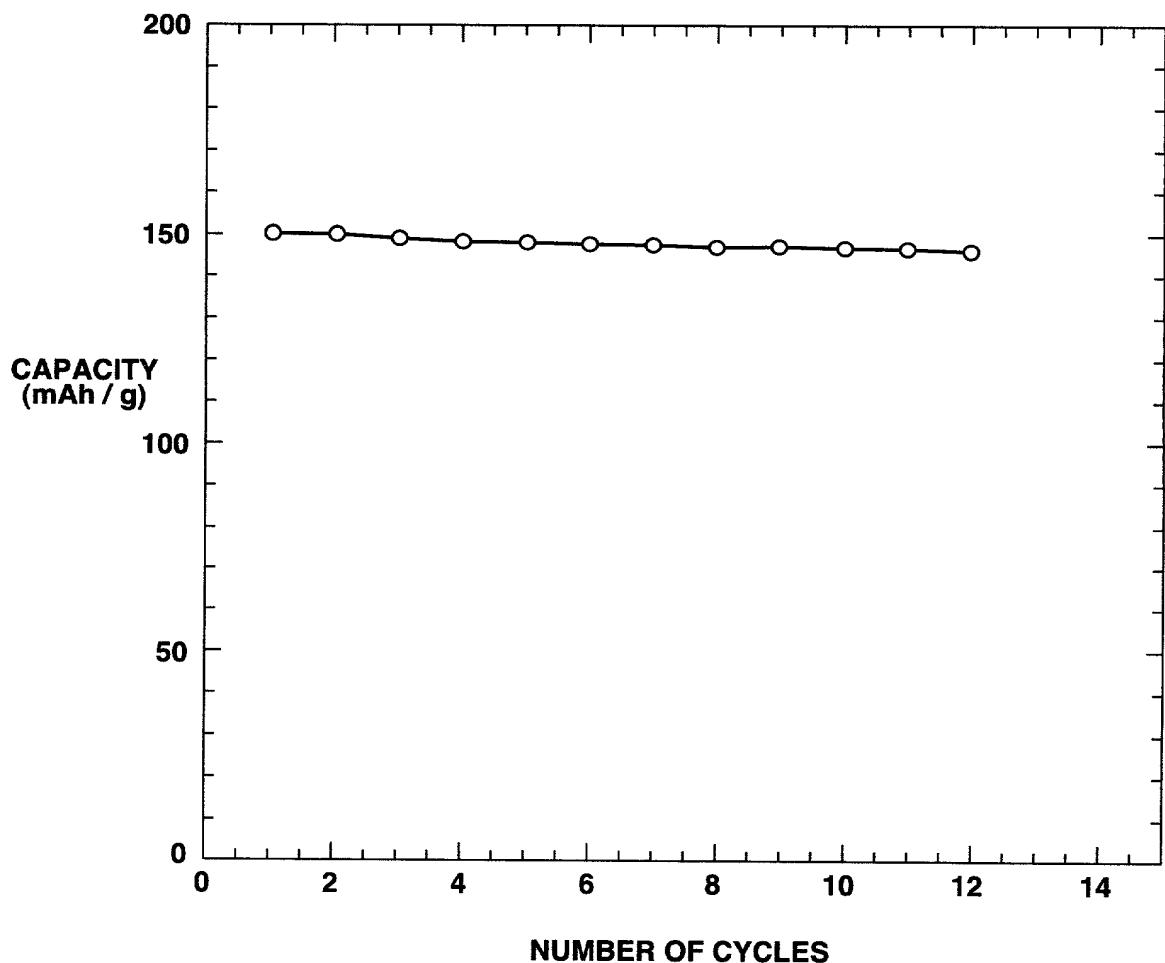


FIG.20

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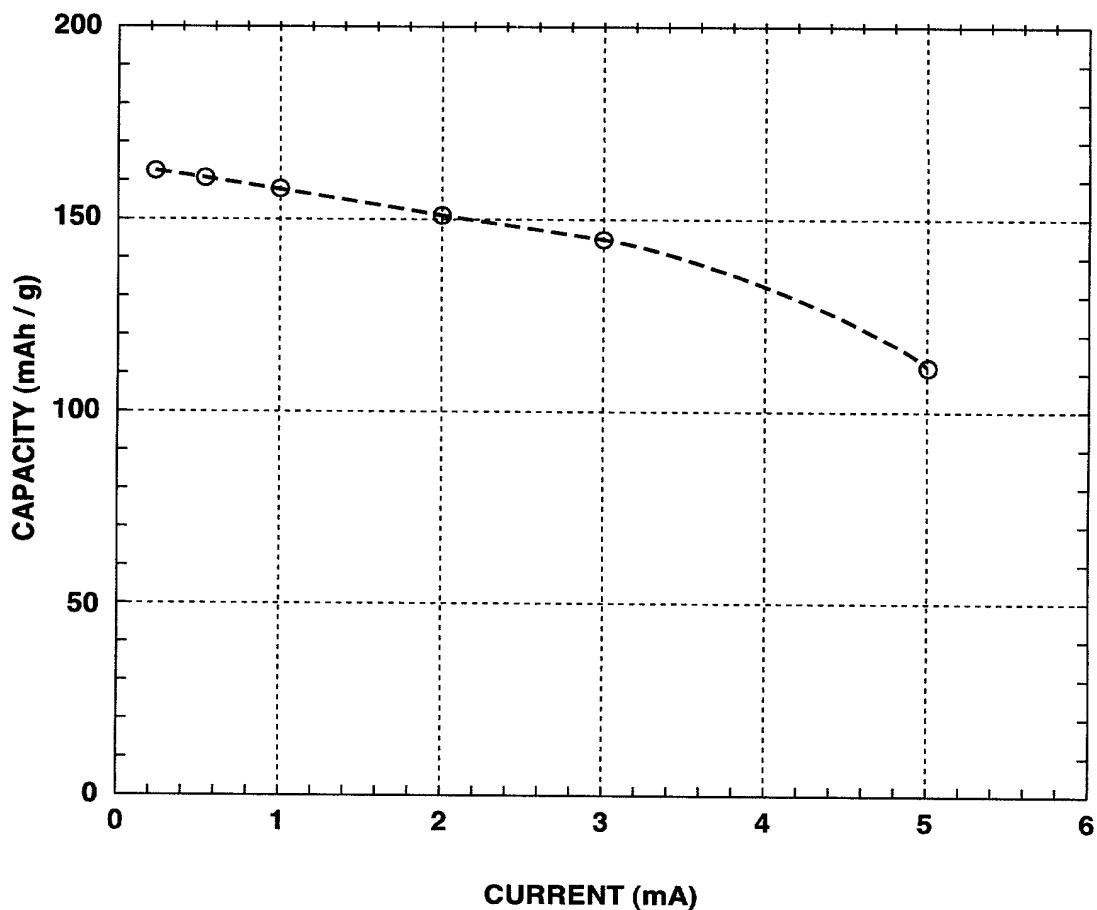


FIG.21

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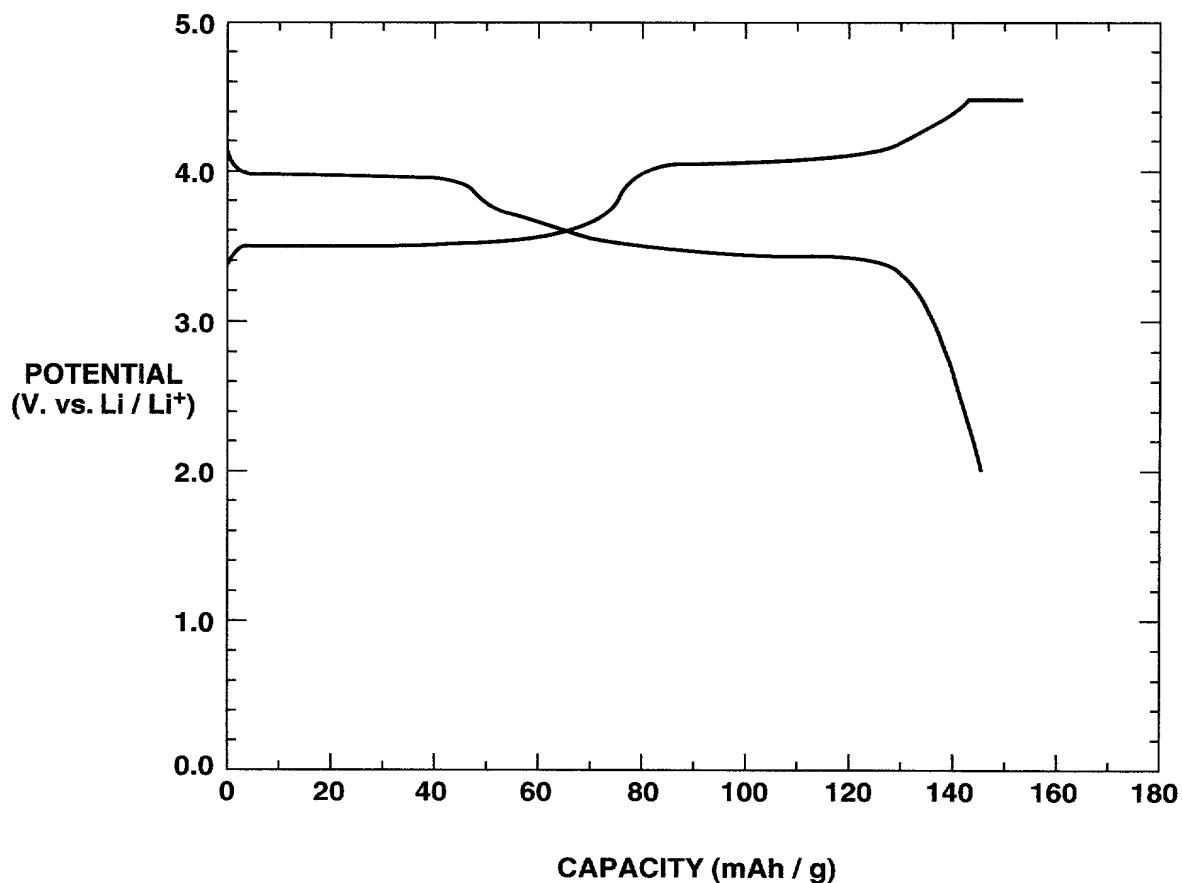


FIG.22

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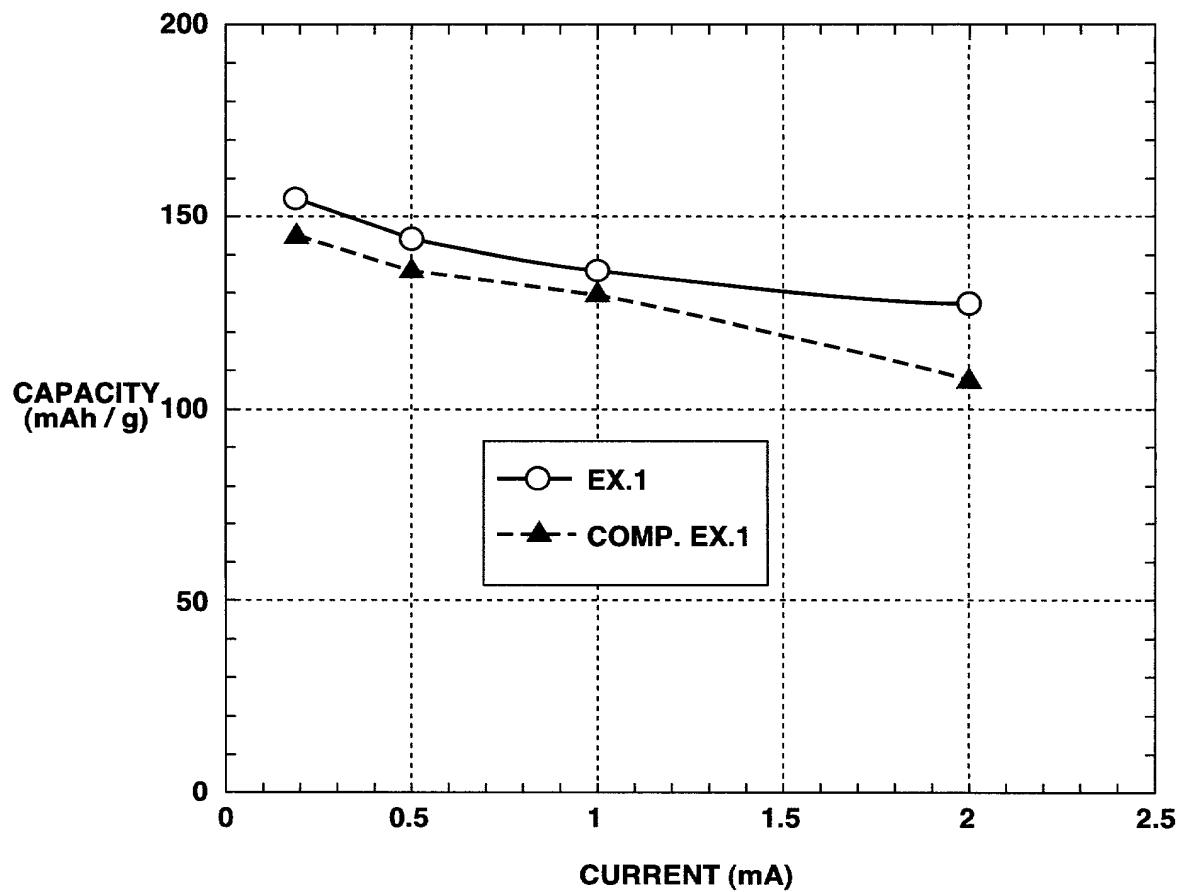


FIG.23

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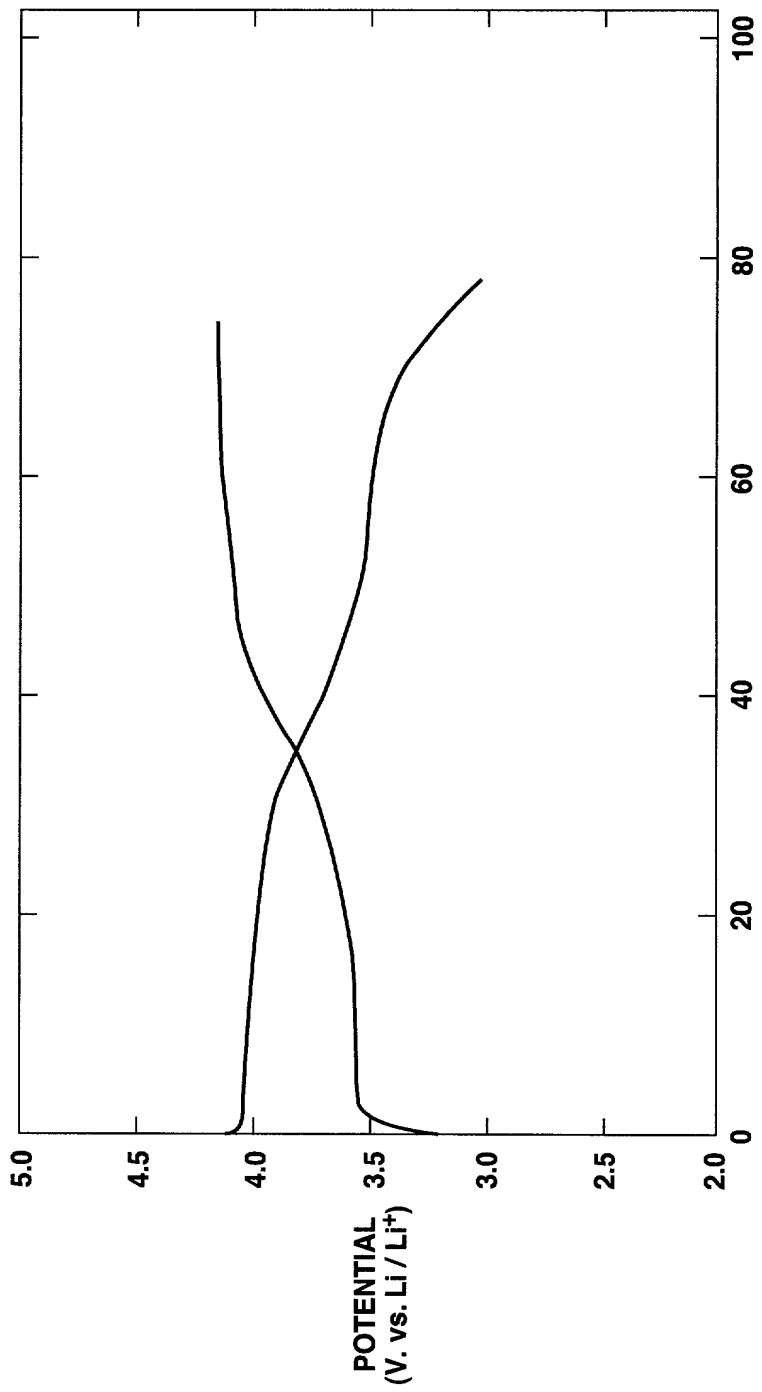


FIG.24

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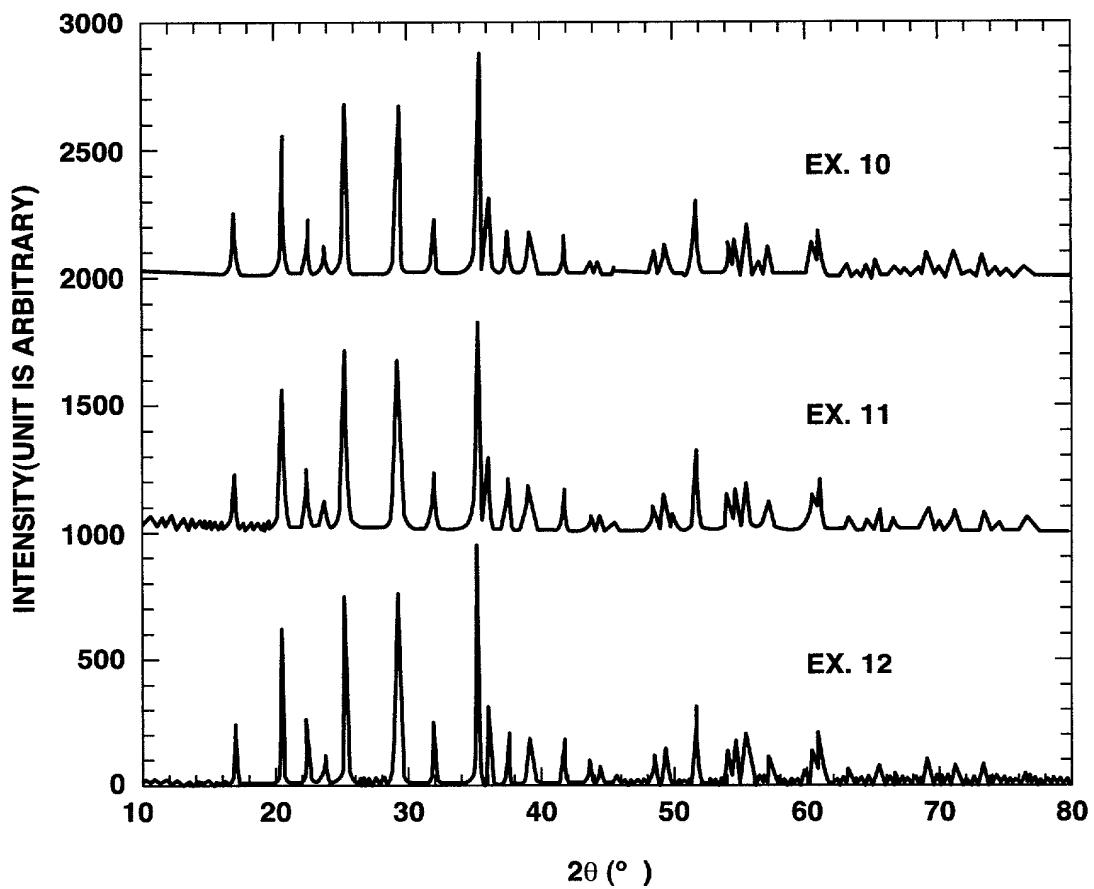


FIG.25

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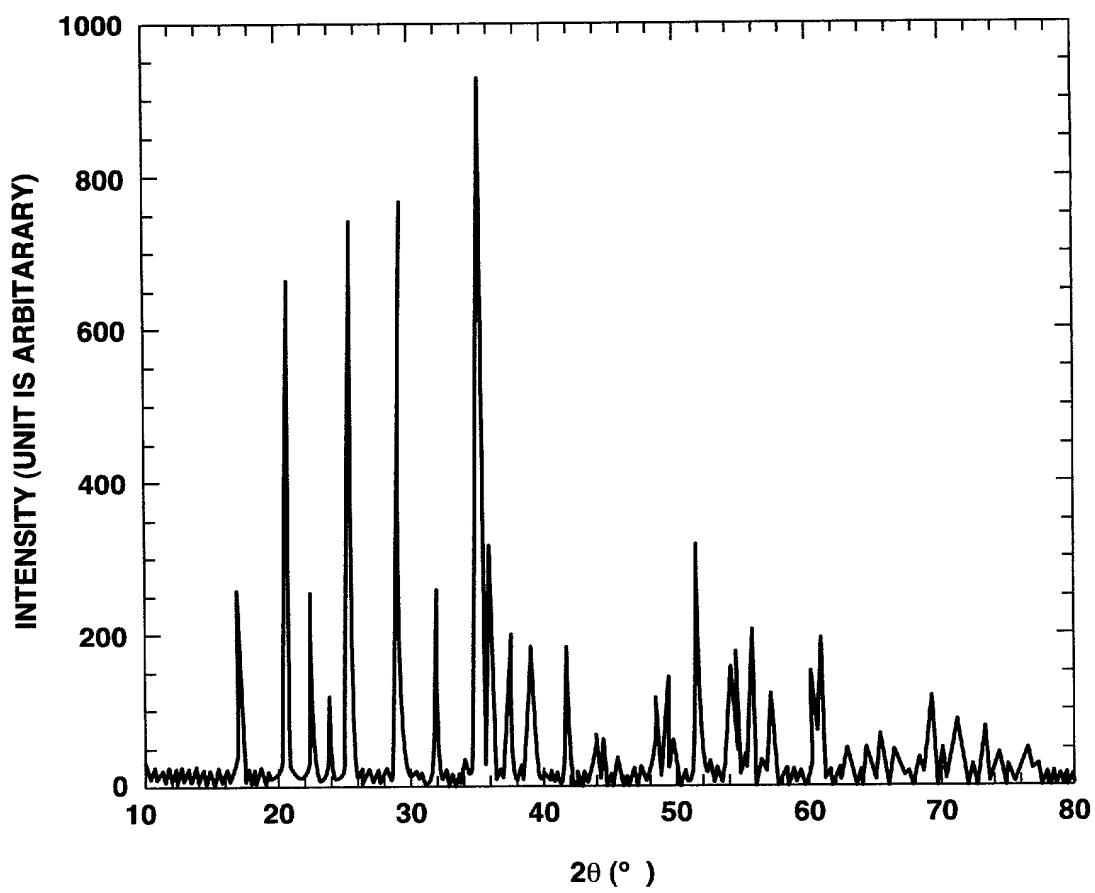


FIG.26

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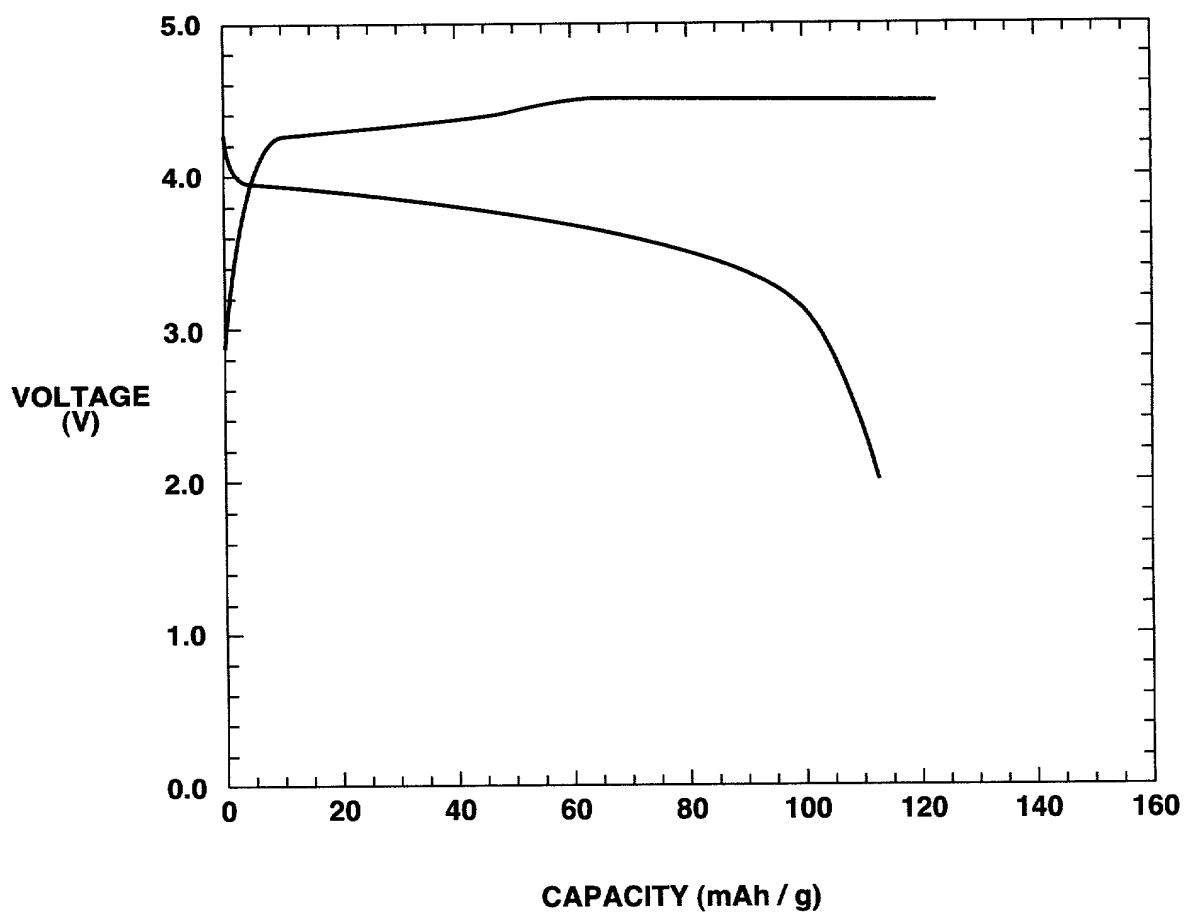


FIG.27

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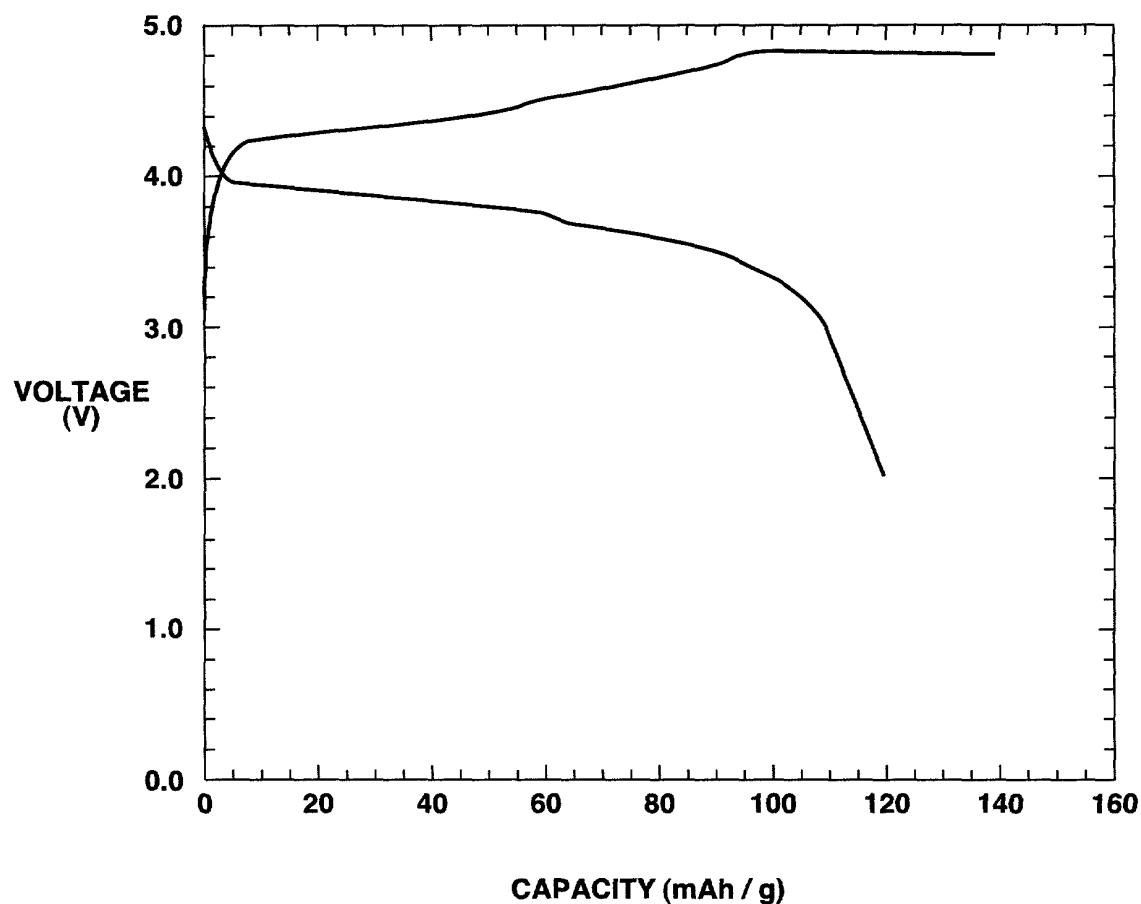


FIG.28

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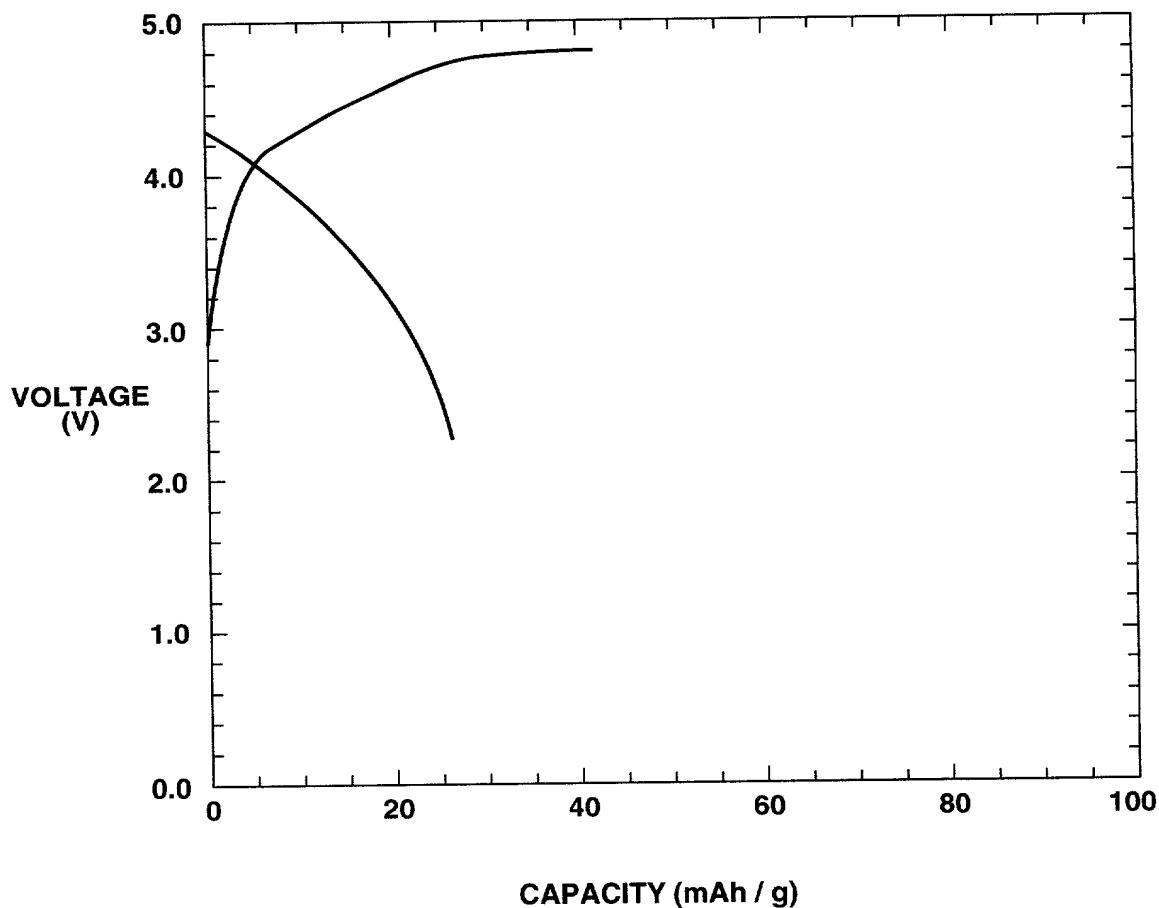


FIG.29

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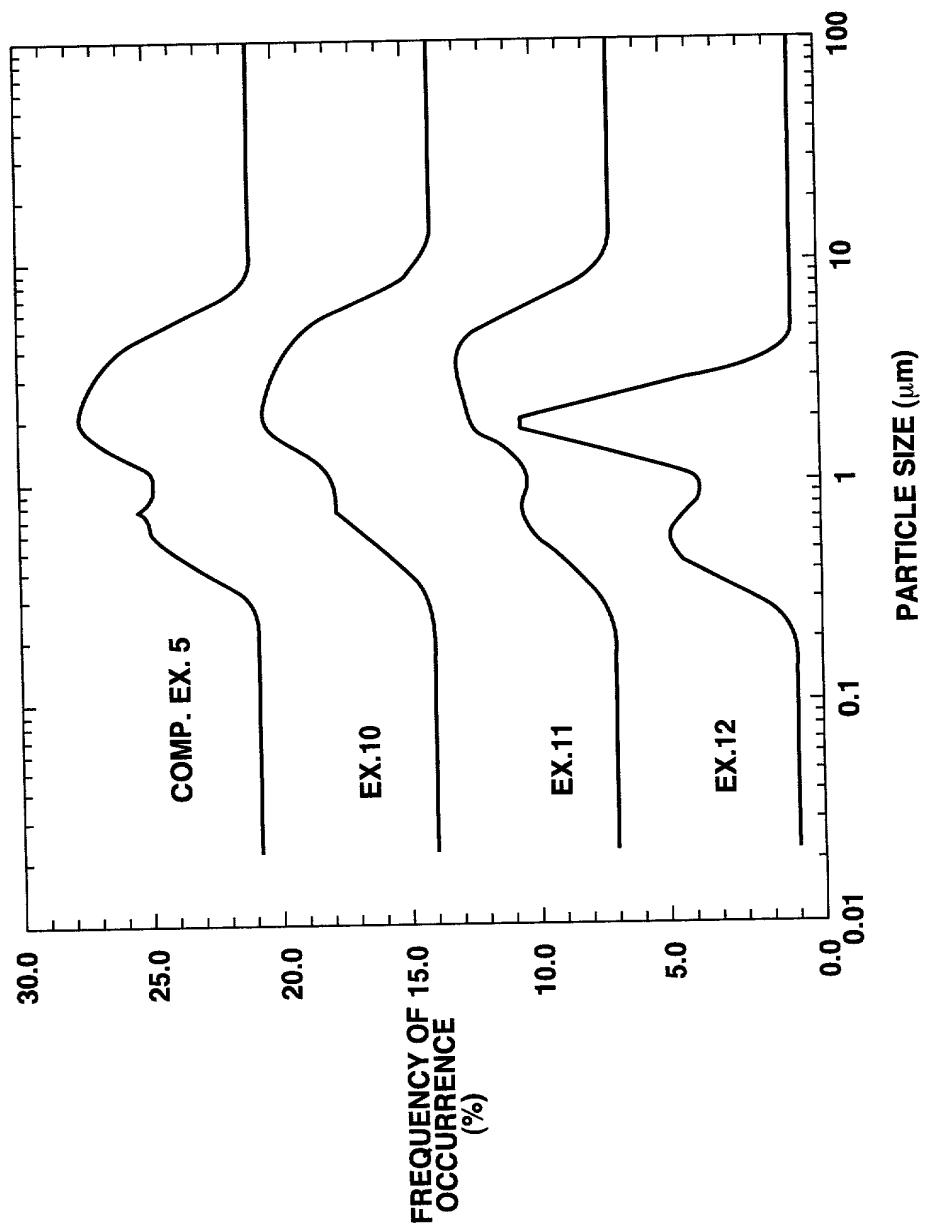


FIG. 30

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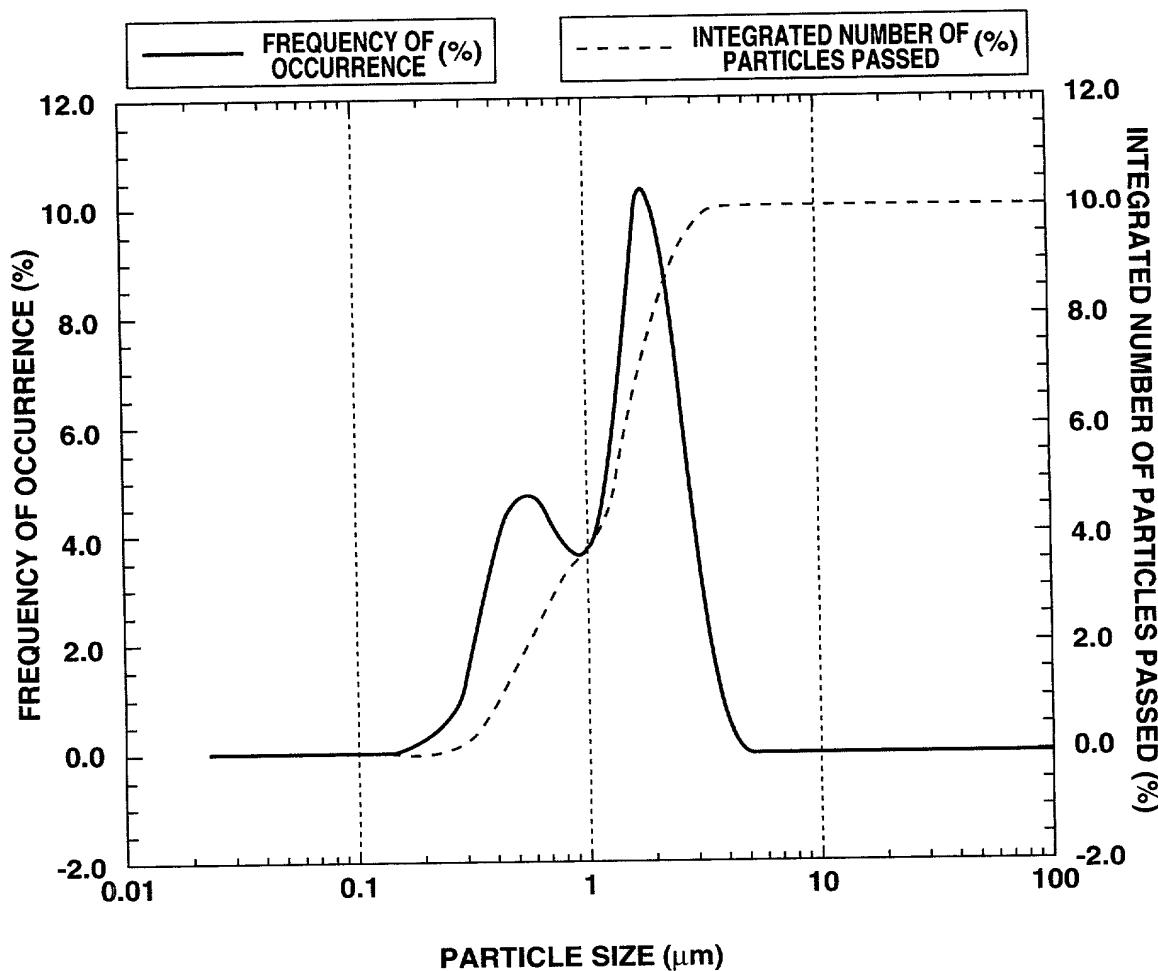


FIG.31

DECLARATION AND POWER OF ATTORNEY FOR PATENT APPLICATION

特許出願宣言書及び委任状

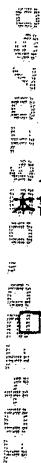
Japanese Language Declaration

日本語宣言書

下記の氏名の発明者として、私は以下の通り宣言します。

私の住所、私書箱、国籍は下記の私の氏名の後に記載された通りです。

下記の名称の発明に関して請求範囲に記載され、特許出願している発明内容について、私が最初かつ唯一の発明者（下記の氏名が一つの場合）もしくは最初かつ共同発明者であると（下記の名称が複数の場合）信じています。



上記発明の明細書（下記の欄でx印）がついていない場合は、
本書に添付）は、

一月一日に提出され、米国出願番号または特許協定条約
国際出願番号を_____とし、
(該当する場合)_____に訂正されました。

私は、特許請求範囲を含む上記訂正後の明細書を検討し、
内容を理解していることをここに表明します。

私は、連邦規則法典第37編第1条56項に定義されると
おり、特許資格の有無について重要な情報を開示する義務が
あることを認めます。

As a below named inventor, I hereby declare that:

My residence, post office address and citizenship are as stated next to my name,

I believe I am the original, first and sole inventor (if only one name is listed below) or an original, first and joint inventor (if plural names are listed below) of the subject matter which is claimed and for which a patent is sought on the invention entitled

METHOD FOR PRODUCING POSITIVE ELECTRODE ACTIVE MATERIAL
AND METHOD FOR PRODUCING NON-AQUEOUS ELECTROLYTE
SECONDARY BATTERY

the specification of which is attached hereto unless the following box is checked:

was filed on March 28, 2000 as United States Application Number or PCT International Application Number PCT/JP00/01915 and was amended on _____ (if applicable)

I hereby state that I have reviewed and understand the contents of the above identified specification, including the claims, as amended by any amendment referred to above.

I acknowledge the duty to disclose information which is material to patentability as defined in Title 37, Code of Federal Regulations, Section 1.56.

私は、米国法典第35編119条(a)-(d)項又は365条(b)項に基づき下記の、米国以外の国の少なくとも一ヵ国を指定している特許協力条約365(a)項に基づく国際出願、又は外国での特許出願もしくは発明者証の出願についての外国優先権をここに主張するとともに、優先権を主張している。本出願の前に出願された特許または発明者証の外国出願を以下に、枠内をマークすることで、示しています。

Prior Foreign Application(s)
外国での先行出願

P11-099407 Japan 06 April 1999
(Number) (Country) (Day Month Year Filed)
(番号) (国名) (出願年月日)

P11-274746 Japan 28 September 1999
(Number) (Country) (Day Month Year Filed)
(番号) (国名) (出願年月日)

P11-274747 Japan 28 September 1999
(Number) (Country) (Day Month Year Filed)
(番号) (国名) (出願年月日)

(Number) (Country) (Day Month Year Filed)
(番号) (国名) (出願年月日)

私は、第35編米国法典119条(e)項に基いて下記の米国特許出願規定に記載された権利をここに主張いたします。

(Application No.) (Filing Date)
[出願番号] [出願日]

I hereby claim foreign priority under Title 35, United States Code, Section 119(a)-(d) or 365(b) of any foreign application(s) for patent or Inventor's certificate or 365(a) of any PCT International application which designated at least one country other than the United States, listed below and have also identified below, by checking the box, any foreign application for patent or Inventor's certificate or PCT International application having a filing date before that of the application on which priority is claimed:

Priority Not Claimed
優先権主張なし

(Number) (Country) (Day Month Year Filed)

I hereby claim the benefit under Title 35, United States Code, Section 119(e) of any United States provisional application(s) listed below.

(Application No.) (Filing Date)
[出願番号] [出願日]

I hereby claim the benefit under Title 35, United States Code, Section 120 of any United States application(s) or 365(c) of an PCT International application designating the United States, listed below and, insofar as the subject matter of each of the claims of this application is not disclosed in the prior United States or PCT International application in the manner provided by the first paragraph of Title 35, United States Code, Section 112, I acknowledge the duty to disclose information which is material to patentability as defined in Title 37, Code of Federal Regulations, Section 1.56 which became available between the filing date of the prior application and the national or PCT International filing date of this application.

(Application No.)
(出願番号)

(Filing Date)
(出願日)

(Status: patented, pending, abandoned)

(現況: 特許許可済、権属中、放棄済)

Application No.)
(出願番号)

(Filing Date)
(出願日)

(Status: patented, pending, abandoned)

(現況: 特許許可済、権属中、放棄済)

私は、私自身の知識に基づいて本宣言書中で私が行なう表明が眞実であり、かつ私の入手した情報と私の信じるところに基づく表明が全て眞実であると信じてのこと、さらに故意になされた虚偽の表明及びそれと同等の行為は米国法典第18編第1001条に基づき、罰金または拘禁、もしくはその両方により処罰されること、そしてそのような故意による虚偽の声明を行なえば、出願した、又は既に許可された特許の有効性が失われることを認識し、よってここに上記のごとく宣誓を致します。

I hereby declare that all statements made herein of my own knowledge are true and that all statements made on information and belief are believed to be true; and further that these statements were made with the knowledge that willful false statements and the like so made are punishable by fine or imprisonment, or both, under Section 1001 of Title 18 of the United States Code and that such willful false statements may jeopardize the validity of the application or any patent issued thereon.

委任状: 私は下記の発明者として、本出願に関する一切の手続きを米特許商標局に対して遂行する弁理士または代理人として、下記の者を指名いたします。(弁護士、または代理人の氏名及び登録番号を明記のこと)

POWER OF ATTORNEY: As a named Inventor, I hereby appoint the following attorney(s) and/or agent(s) to prosecute this application and transact all business in the Patent and Trademark Office connected therewith. (list name and registration number)

David R. Metzger (Reg. No. 32,949), Joseph A. Mahoney (Reg. No. 38,956), Howard B. Rockman (Reg. No. 22,190), Jordan A. Sigale, (Reg. No. 39,028), Michael A. Molano (Reg. No. 39,777), Michael L. Kiklis (Reg. No. 38,939), Janelle D. Strode (Reg. No. 34,738), Kevin W. Guynn (Reg. No. 29,972), Jennifer Hammond (Reg. No. 41,814), Lana Knedlik (Reg. No. 42,748), John F. Griffith (Reg. No. 44,137), Marina Saito (Reg. No. 42,121), Alison P. Schwartz (Reg. No. 43,863), Christopher P. Rauch (Reg. No. 45,034), Francisco Rubio-Campos (Reg. No. 45,358), Brian J. Gill (Reg. No. 46,727) and Shashank S. Upadhye, all members of the firm of Sonnenschein, Nath & Rosenthal

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書類送付先

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Chicago, Illinois 60606-1080

直接電話連絡先: (名前及び電話番号)

Direct Telephone Calls to: (name and telephone number)

312/876-2578

唯一または第一発明者名		Full name of sole or first inventor: Guohua Li
発明者の署名	日付	Inventor's signature Guohua Li Date Dec. 22, 2000
住所	Residence Kanagawa, Japan	
国籍	Citizenship China	
私書箱	Post Office Address c/o Sony Corporation, 7-35, Kitashinagawa 6-chome, Shinagawa-ku, Tokyo 141-0001 Japan <i>JPX</i>	

(第三以降の共同発明者についても同様に記載し、署名をすること)

第二共同発明者		Full name of second joint inventor, if any: Atsuo YAMADA
発明者の署名	日付	Inventor's signature <i>Atsuo Yamada</i> Date Dec. 27, 2000
住所	Residence Kanagawa, Japan	
国籍	Citizenship Japan <i>JPX</i>	
私書箱	Post Office Address c/o Sony Corporation, 7-35, Kitashinagawa 6-chome, Shinagawa-ku, Tokyo 141-0001 Japan	
第三共同発明者		Full name of third joint inventor, if any: Hideto AZUMA
発明者の署名	日付	Inventor's signature <i>Hideto Azuma</i> Date Jan. 15, 2001
住所	Residence Kanagawa, Japan	
国籍	Citizenship Japan <i>JPX</i>	
私書箱	Post Office Address c/o Sony Corporation, 7-35, Kitashinagawa 6-chome, Shinagawa-ku, Tokyo 141-0001 Japan	

Japanese Language Declaration

(日本語宣言書)